

IN VITRO ANALYSIS OF ANTIOXIDANT ACTIVITIES OF *OXALIS CORNICULATA* LINN.  
FRACTIONS IN VARIOUS SOLVENTS

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### Abstract

As part of our search for natural antioxidants, this work presents an evaluation of antioxidant activities of methanolic extract of *Oxalis corniculata* and its sub-fractions in hexane, chloroform, ethyl acetate, *n*-butanol and water. The total phenolic contents in terms of  $\mu\text{g}$  of gallic acid equivalents per mg of dried mass were approximately 21.0, 28.2, 34.5, 162.0, 70.0, and 49.2 in methanolic, hexane, chloroform, ethyl acetate, *n*-butanolic and aqueous fractions respectively, while the flavonoid contents in these solvents were 362.4, 214.1, 317.1, 177.1, 98.8 and 53.5 respectively in terms of  $\mu\text{g}$  of rutin per mg of dried mass. In DPPH assay, the ethyl acetate fraction showed the highest free radical scavenging activity, 24.0% with 1 mg/mL concentration. The second strongest fraction was chloroform (21.5%). The  $\text{EC}_{50}$  and  $\text{TEC}_{50}$  values of the methanolic extract were 3.63 mg/mL and 23 min respectively. The FRAP values in terms of  $\mu\text{g}$  of ascorbic acid equivalents per mg of dried mass for these solvents were 288.0, 1705.3, 437.1, 72.0, 28.0, and 44.0 respectively while total antioxidant activity measured by phosphomolybdate assay in terms of  $\mu\text{g}$  of ascorbic acid equivalents per mg of dried mass were 50.0, 117.0, 78.6, 57.8, 3.4 and 8.3 respectively. All the samples showed remarkable ability to inhibit lipid peroxidation exhibiting much better and sustainable peroxidation inhibitory activity than the standard butylated hydroxyanisole.

**Key words:** *Oxalis corniculata*, antioxidant, free radical, phenolics, flavonoids

### Introduction

During the metabolic process in human body a number of free radicals are formed which can cause harmful effects leading to various diseases, such as, cardiovascular ailments, inflammation, cancer, neural disorders, Alzheimer disease, arteriosclerosis, allergies, and aging (Tiganis, 2011; Oktay and Gulcin, 2003; Niki et al., 1994; Cook and Samman, 1996; Sohal, 2002; Choi et al., 2002; Finkel and Holbrook, 2000). To control these reactive free radicals, a number of synthetic antioxidants are available but they have harmful side effects (Rechner et al., 2002). There is, thus, a dire need of natural antioxidants which are not only safer but also cheaper and more easily available to a common person (Madsen, 1995; Lim and Murtiyaja, 2007). Plant kingdom that possesses a virtually unending reservoir of therapeutic and nutraceutical agents must, therefore, be explored for chemical principles having antioxidant properties (Mishra et al., 2007).

*Oxalis corniculata* (Oxalidaceae) is a cosmopolitan herb and abundantly found in agricultural farms, gardens and lawns, etc. (Lourteig, 2000; Maberley, 2008). In Pakistan, the plant is distributed through out the country (Nasir, 1971). It is a small perennial herb with leaves having sour taste, and locally known as *khatcorla* or *khati buti* in Hindko (Shah and Khan, 2006b). In folklore, the plant is used to cure a number of diseases such as indigestion and stomach disorders, liver and bladder inflammation and jaundice (Qureshi et al., 2008; Hussain et al. 2008; Abbasi et al., 2010). Fresh leaves are also applied as antiscorbutic (Shah and Khan, 2006b), aperient, diaphoretic, diuretic, antidiabetic, enthelminthic (Shah and Khan, 2006a). The crushed fresh leaves are used to stop bleeding from wounds (Hebbar, et al., 2004).

*Oxalis corniculata* has also been reviewed (Badwaik et al., 2011; Kathiriya et al., 2010) highlighting its diverse ethnomedicinal applications and research studies carried out on the plant so far. The plant has been reported to contain flavonoids, tannins, phytosterol, phenolics, glycosides, fatty acids, volatile oil, vitamin C and oxalates (Badwaik et al., 2011). Sakat et al. (2009) evaluated phenolics and flavonoids contents of the defatted methanolic extract of *Oxalis corniculata* from Mumbai, India, and analyzed the antioxidant properties of the extract using DPPH and nitric oxide free radical scavenging assays. Reddy et al. (2010) also investigated the antioxidant properties of the methanolic extract of the plant from Talakona (India) using DPPH assay and reducing power method. More recently, Alam et al. (2011) conducted a more detailed evaluation of the antioxidant activities of the plant from Khulna, Bangladesh. The study was conducted on crude methanolic extract and its sub-fractions in dichloromethane, ethyl acetate, and *n*-butanol to evaluate phenolic and flavonoid contents, total antioxidant activity, DPPH free radical scavenging, nitric oxide radical scavenging, and reducing power activity. In these studies, the crude methanolic extract was obtained through Soxhlet extraction.

In view of its extensive ethnomedicinal application by the local people, we investigated the antioxidant properties of the Pakistani variety of *Oxalis corniculata*. In the study, a different methodology was used to obtain the crude methanolic extract and its fractions in various solvents.

## Materials and Methods

### Chemicals

Ascorbic acid and DPPH [1,1-diphenyl-2-picrylhydrazyl] radical were purchased from MP biomedical, Inc. (France). Rutin, Linoleic acid and Folin-Ciocalteu reagent were purchased from Sigma Aldrich. Butylated hydroxyanisole (BHA) and Tween-20 were obtained from Merck (Darmstadt, Germany). Gallic acid was purchased from Scharlau-Switzerland and 2,4,6-(2-tripyrilydyl)-S-triazine (TPTZ) was obtained from Alfa-Aesar, (Germany). All other solvents and chemicals used were of analytical grade.

### Plant material and Preparation of plant extract and fractions

The aerial parts of the herb, *Oxalis corniculata*, were collected from the hilly area near Abbottabad, Hazara (Pakistan) in June 2010 and identified by the taxonomist. The plant was cleaned and air-dried in the shade for ten days and then ground to obtain a fine powder. The powdered material (100 g) was soaked in 500 mL of methanol and kept at room temperature for ten days. The extract was filtered and the solvent was evaporated from the filtrate under reduced pressure with rotary evaporator to obtain crude methanolic (MeOH) extract. The MeOH extract (1.90 g) was suspended in double distilled water and extracted successively with hexane, chloroform (CHCl<sub>3</sub>), ethyl acetate (EtOAc) and *n*-butanol (*n*-BuOH). In this way, six samples were obtained, namely, MeOH extract, and hexane, CHCl<sub>3</sub>, EtOAc, *n*-BuOH and aqueous (H<sub>2</sub>O) fractions. They were dried under reduced pressure, weighed and kept in refrigerator until further use.

### Total phenolic content

The total phenolic content in the crude methanolic extract and its sub-fractions of *O. corniculata* was determined according to a well-cited protocol (Singleton and Rossi, 1965). Briefly, 1.5 mg of an extract was dissolved in 5 mL of methanol from which 40 µL was taken and dissolved in 3.16 mL of distilled water. To this 200 µL of Folin-Ciocalteu reagent was mixed and, after an interval of 8 min, 20% of 600 µL sodium carbonate solution was added. The mixture was incubated at 40 °C for 30 min and absorbance was measured at 765 nm on UV/Visible spectrophotometer. The standard calibration curve was prepared with gallic acid standard solution (50 to 500 mg/L) following the same procedure. The gallic acid equivalent (GAE) was determined from the following equation of the standard curve and the results were expressed as µg of GAE per mg of dried sample.

$$\text{GAE} = [\text{Absorbance (765 nm)} / 0.0009567] + 0.02091, R^2 0.9995.$$

### Total flavonoid content

The total flavonoid contents in the samples were determined following the method reported by Park *et al.* (2008). In each experiment, 1.5 mg of methanolic extract or a fraction was dissolved in 5 mL of methanol, from which 300 µL were transferred into 3.4 mL of 30% aqueous methanol. To this mixture, 150 µL each of 0.5M NaNO<sub>2</sub> and 0.3 M AlCl<sub>3</sub>.6H<sub>2</sub>O were added and mixed thoroughly. After the interval of 5 min, 1 mL of 1 M NaOH was also added. The absorbance was observed immediately at 506 nm against a blank. The standard curve was plotted using rutin as a standard following the same method. The results were expressed as µg of Rutin per mg of dried mass as obtained from the following equation.

$$\text{Rutin Equivalent (RE)} = \text{Absorbance (506 nm)} / 0.0002428 - 0.008141, R^2 0.9987.$$

### Free radical scavenging activity by DPPH assay

The free radical scavenging activities of the samples by DPPH method were determined according to the method reported by Brand-Williams *et al.* (1995). In this method, 2.4 mg of DPPH free radical were dissolved in 100 mL of methanol to prepare its stock solution, which was kept at 20 °C until required. The working solution of DPPH was obtained by diluting its stock solution with methanol till the absorbance was noted to be 0.980 ± 0.02 at 517 nm. Then, 3 mL of the working solution was mixed with 100 µL of a sample (1 mg/mL). After incubating the mixture in the dark for 30 min, absorbance was measured at 517 nm. The scavenging activity was calculated by using the formula:

$$\% \text{ inhibition} = [(\text{Absorbance of blank} - \text{Absorbance of sample}) / \text{Absorbance of blank}] \times 100$$

The blank contained all reagents except the test sample. In order to establish the relationship between concentration of methanolic extract and its free radical scavenging activity, solutions of extract with different concentrations were prepared and their activities were studied. The EC<sub>50</sub> value was also determined. To determine the variation in free radical scavenging activity of the extract with time, %DPPH remaining was also found out using the following formula:

$$\% \text{DPPH}_{\text{rem}} = \frac{[\text{DPPH}]_{T=t}}{[\text{DPPH}]_{T=0}}$$

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Where  $[DPPH]_{T=0}$  is the concentration of DPPH radical before reaction with antioxidant sample and  $[DPPH]_{T=t}$  is the concentration of DPPH radical after reaction with antioxidant sample at any time  $t$ . From the data, TEC50 was also calculated which is the time required for 50% inhibition of DPPH by the extract.

#### Reducing power by FRAP assay

The ferric reducing antioxidant power (FRAP) assay was carried out according to the protocol reported by Benzie and Strain (1999). The FRAP reagent is prepared as follows: 300 mM sodium acetate buffer (pH 3.6, 25 mL), 10 mM TPTZ solution (2.5 mL) in 40 mM HCl solution (20 mL) and 20 mM  $FeCl_3 \cdot 6H_2O$  solution (2.5 mL) were mixed thoroughly and kept at 37 °C throughout the experiment. After initial incubation for 10 min, the absorbance of the reagent was noted at 593 nm. For a test, 3 mL of FRAP reagent was added to 100  $\mu$ L of sample (1 mg/mL) and absorbance was noted at 593 nm after 5 min. The blank was prepared by dissolving 100  $\mu$ L of methanol in 3 mL of FRAP reagent. Ascorbic acid was used as a standard antioxidant and the same procedure was employed to determine its FRAP values for different concentrations and a calibration curve was plotted. The results were expressed in terms of  $\mu$ g of ascorbic acid equivalents (AAE) per mg of the dried mass, which was determined from the following equation;

$$\text{Ascorbic Acid Equivalent} = [(\text{Absorbance at } 593 \text{ nm} / 0.002) - 0.004], R^2 = 1$$

#### Total antioxidant activity by phosphomolybdate assay

To carry out phosphomolybdate assay, the procedure followed by Umamaheswari and Chatterjee (2008) was used. The phosphomolybdate reagent was prepared by mixing equal volumes (100 mL) of 0.6 M sulfuric acid, 28 mM sodium phosphate and 4 mM ammonium molybdate. Test samples were prepared by dissolving 1 mg of plant methanolic extract or any of its sub-fraction in 1 mL of methanol. Then, 0.1 mL of the sample was dissolved in 1 mL of reagent solution in a test tube which was capped with silver foil and incubated in water bath at 95 °C for 90 min. After cooling the sample to room temperature, the absorbance was observed at 765 nm against a blank. Ascorbic acid was used as a standard antioxidant. Various concentrations of ascorbic acid (50 to 500 mg/L) were prepared and the same protocol was carried out in order to plot a standard curve. The results were expressed as  $\mu$ g of ascorbic acid equivalent (AAE) per mg of the dried weight of the sample as determined from the equation of the standard calibration curve.

$$\text{AAE} = [(\text{Absorbance } (765 \text{ nm}) / 0.0025) + 0.0593], R^2 = 0.997$$

#### Lipid peroxidation inhibition

The lipid peroxidation assay was carried out following the protocol given by Mitsuda *et al.* (1966). To prepare an emulsion of linoleic acid, 155  $\mu$ L of the acid were mixed with 175  $\mu$ g of Tween-20, and the mixture was put into 50 mL of potassium phosphate buffer (pH 7). Meanwhile, 5 mg of a sample was dissolved in 1 mL of methanol. An aliquot of 100  $\mu$ L was dissolved in 2.4 mL of potassium phosphate buffer (0.04 M, pH 7) and the mixture was put in 2.5 mL of linoleic acid emulsion. The mixture was incubated at 37 °C. The 100  $\mu$ L of aliquot from the incubated solution was regularly taken at 24 h intervals and allowed to react with 100  $\mu$ L each of 20 mM  $FeCl_2$  and 30% ammonium thiocyanate. The absorbance was noted at 500 nm after every 24 h interval for 4 days. A 5 mL solution consisting of equal quantities of linoleic acid emulsion and potassium phosphate buffer was taken as a blank. Butylated hydroxyanisole (BHA) was used as a standard and same procedure was followed to measure its ability to inhibit peroxidation in linoleic acid.

#### Statistical analysis

All experiments were carried out in triplicate and mean values were calculated. One way analysis of variance (ANOVA) was applied and the results were correlated.

## Results and Discussion

### Total phenolic content

The methanolic extract and fractions in various solvents of *Oxalis corniculata* were subjected to evaluation of total phenolic content and the results are shown in Table 1. In this study on the Pakistani variety of the plant, ethyl acetate (EtOAc) fraction showed the highest phenolic content (162.03  $\mu$ g of Gallic Acid Equivalent per mg of the dried mass, GAE), while the crude methanolic (MeOH) extract had the lowest value (20.91 GAE). The MeOH value was close to that reported earlier by Sakat *et al.* (2010) for the Indian plant. The phenolic content in various solvents decreases in the order of EtOAc > *n*-BuOH > H<sub>2</sub>O, the trend also exhibited by the Bangladeshi variety (Alam *et al.*, 2011). However, the crude methanolic extract of Bangladeshi plant showed a higher value of total phenolics than both the Pakistani and Indian. The reason may lie in the methods of extraction as well as climatic conditions.

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**Table 1:** Percent yield, total flavonoid and total phenolic contents of *Oxalis corniculata*

Extract/Fraction	Percent yield*	Total flavonoids in $\mu\text{g}$ of **RE /mg of dried mass	Total phenolics in $\mu\text{g}$ of $^{\dagger}$ GAE /mg of dried mass
MeOH	---	362.43 $\pm$ 9.26	20.91 $\pm$ 4.26
Hexane	29.73	214.14 $\pm$ 3.31	28.23 $\pm$ 2.08
CHCl <sub>3</sub>	22.12	317.13 $\pm$ 7.50	34.50 $\pm$ 2.62
EtOAc	16.24	177.09 $\pm$ 5.89	162.03 $\pm$ 9.98
n-BuOH	14.51	98.82 $\pm$ 2.76	70.05 $\pm$ 7.30
H <sub>2</sub> O	15.40	53.52 $\pm$ 1.32	49.15 $\pm$ 4.45

\*The amount of crude methanolic extract was 1.90 g; \*\*Rutin equivalents;  $^{\dagger}$ Gallic acid equivalents; Each value listed in the table is expressed as mean  $\pm$  SD (n=3).

**Table 2:** Free radical scavenging by DPPH assay, reducing power by FRAP assay, total antioxidant activity by phosphomolybdate assay

Extract / Fraction	%Free radical scavenging by DPPH assay*	Reducing power by FRAP assay $\mu\text{g}$ of **AAE/ mg of dried mass	Total antioxidant activity by phosphomolybdate assay $\mu\text{g}$ of **AAE/mg of dried mass
MeOH	3.67	287.98 $\pm$ 5.26	49.80 $\pm$ 2.91
Hexane	1.68	1705.26 $\pm$ 9.57	117.00 $\pm$ 7.26
CHCl <sub>3</sub>	21.45	437.13 $\pm$ 3.24	78.60 $\pm$ 2.20
EtOAc	24.03	71.98 $\pm$ 3.33	57.80 $\pm$ 8.10
n-BuOH	7.54	27.96 $\pm$ 0.98	3.40 $\pm$ 0.60
H <sub>2</sub> O	0.00	43.85 $\pm$ 1.27	8.30 $\pm$ 0.72

\*The amount of crude methanolic extract was 1.90 g; \*\*Rutin equivalents;  $^{\dagger}$ Gallic acid equivalents; Each value listed in the table is expressed as mean  $\pm$  SD (n=3).

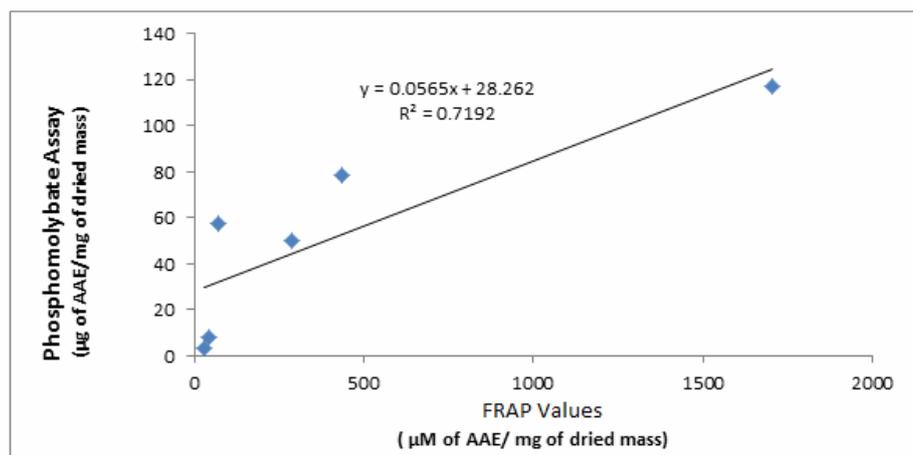
**Table 3:** Change of % inhibition of DPPH free radical with changing concentration of methanolic extract of *Oxalis corniculata*

S.No	Concentration (mg/mL)	% Inhibition
1	0.2	01.79
2	0.4	05.84
3	0.5	08.31
4	0.6	12.35
5	0.7	15.06
6	0.8	17.42
7	0.9	21.66
8	01	25.90
9	02	31.78
10	03	40.60
11	04	53.45
12	05	62.09
1	07	74.71
14.	10	85.57
15	15	98.64
EC <sub>50</sub> (mg/mL)	3.63	

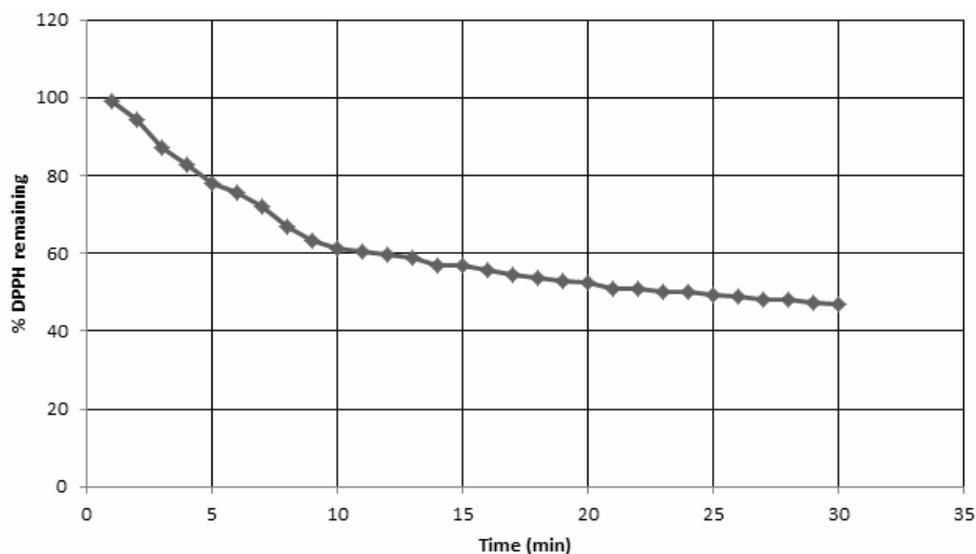
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### Total flavonoid content

Total flavonoid content of the methanolic extract and fractions in various solvents was determined in terms of  $\mu\text{g}$  of Rutin Equivalent per mg of the dried material (RE) and the results are shown in Table 1. In our study, the total flavonoid content decreases in the following order: MeOH >  $\text{CHCl}_3$  > Hexane > EtOAc > n-BuOH >  $\text{H}_2\text{O}$ . The overall trend was the same as reported by Alam *et al.* (2011), however, MeOH extract in our study exhibited a much higher value. In EtOAc, n-BuOH and  $\text{H}_2\text{O}$  fractions, the phenolic and flavonoid contents showed the same trend; EtOAc being the highest and  $\text{H}_2\text{O}$  the lowest. The  $\text{CHCl}_3$  fraction, however, showed much higher value of flavonoids (317.13 RE) than phenolics (34.50 GAE)



**Figure 1:** The significant correlation was observed between the antioxidant activities as determined by phosphomolybdate assay and FRAP assay.

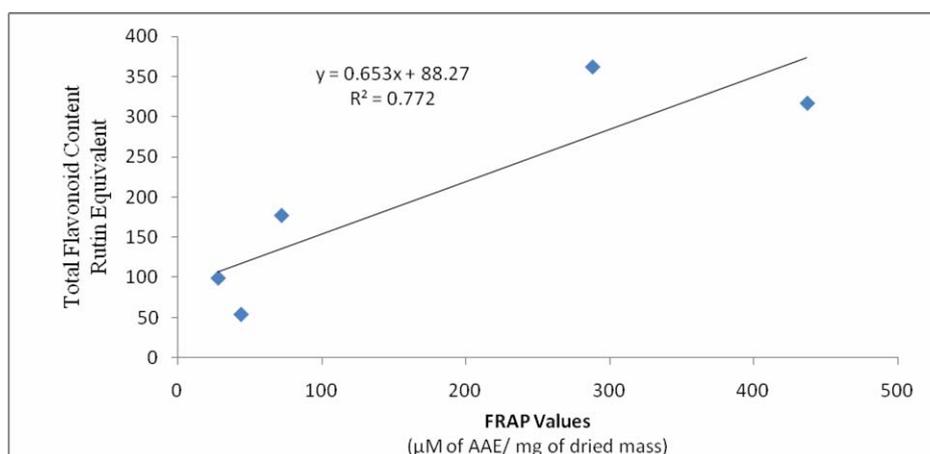


**Figure 2.** Effect of the methanolic extract of *Oxalis corniculata* on DPPH with time. %DPPH remaining decreases with time gradually reaching almost a steady state after 15 min.

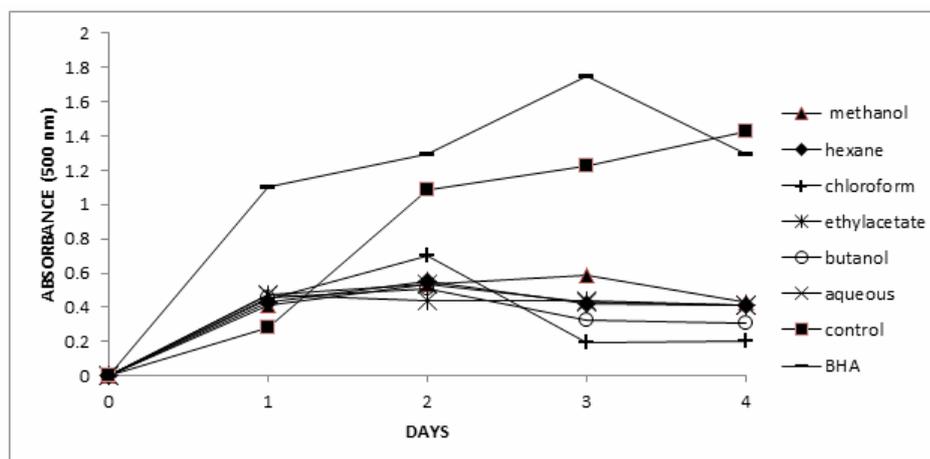
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### Free radical scavenging activity by DPPH assay

Free radical scavenging activities of the samples were determined using well-known DPPH assay and the results are displayed in Table 2. The DPPH is a stable free radical (DPPH<sup>•</sup>) that accepts an electron or a hydrogen radical to form diamagnetic molecule (Oktay et al., 2005; Nakayama, 1994). During the process, the purple colour of the radical turns pale yellow. The DPPH radical shows maximum absorbance at 517 nm. In the present study, the EtOAc fraction exhibited the highest capability to scavenge free radicals. This is in agreement with the data reported by Alam *et al.* (2011). While the CHCl<sub>3</sub> fraction showed good antioxidant potential, the residual H<sub>2</sub>O fraction was completely inactive. The high free radical scavenging activity of EtOAc and CHCl<sub>3</sub> fractions may be related to their high phenolic contents. This is in consonance with findings of other workers as well (Ahmed et al., 2010; Rahman *et al.*, 2008). The hexane fraction proved to be a very weak free radical scavenger, which can be explained on the basis of its very low phenolic content (Table 1). The %DPPH<sub>remaining</sub> decreases with time, first rapidly but then slowly. The TEC<sub>50</sub> value was 23 min which is the time required for 50% decrease in the concentration of DPPH radical. The effect of the change in concentration of the methanolic extract on DPPH radical was also determined, and there was an almost direct relationship between the concentration of the extract and its ability to scavenge radicals. The EC<sub>50</sub> value was 3.63 mg/mL which is the concentration of the extract required to give 50% inhibition of free radical activity of DPPH (Baig et al., 2012).



**Figure 3:** The significant correlation was observed between the total flavonoid content and FRAP values expressed as rutin equivalent and ascorbic acid equivalent respectively.



**Figure 4:** Lipid peroxidation assay. The change in absorbance of ferric thiocyanate in the presence of various samples of *Oxalis corniculata* or BHA over 4 days. \*Each value is the mean of triplicate determinations. Lower absorbance means a stronger antioxidant.

### Reducing power by FRAP assay

In the Ferric Reducing Antioxidant Power (FRAP) assay, ferric ions are reduced to ferrous ions in the presence of an antioxidant (or, a reducing agent) and a coloured ferrous tripyridyltriazine complex is formed at pH 3.6 which is monitored spectrophotometrically ( $\lambda_{\text{max}}$  593 nm) (Huang et al., 2005). The results of the FRAP assay in terms of Ascorbic Acid Equivalent (AAE) are shown in Table 2. The hexane fraction showed extraordinarily high reducing power (1705 AAE). The ability of the samples to reduce iron(III) to iron(II) decreases in the order of hexane > CHCl<sub>3</sub> > MeOH > EtOAc > H<sub>2</sub>O > *n*-BuOH. The FRAP values of various samples show parallel behavior with the flavonoid content (Figure 3), explaining the possible role of flavonoids as potential antioxidants.

### Total antioxidant activity by phosphomolybdate assay

This assay is based on the reduction of molybdenum (VI) to molybdenum(V) which takes place in the presence of a reducing agent (antioxidant). The product is a green phosphomolybdate(V) complex whose formation is monitored with a spectrophotometer. The assay is often used to estimate total antioxidant activity of a sample and the results are expressed in terms of Ascorbic Acid Equivalents (AAE). The total antioxidant activity of various samples decreases in the following order: hexane > CHCl<sub>3</sub> > EtOAc > MeOH > H<sub>2</sub>O > *n*-BuOH. There is a good correlation between the FRAP values and total antioxidant activity measured by phosphomolybdate assay (Figure 1). Here EtOAc fraction showed a higher value than *n*-BuOH fraction which is in agreement with the findings of Alam *et al.* (2011). In our study, the MeOH fraction showed much less antioxidant activity than the findings of Alam *et al.* (2011). The reason possibly lies in the method of extraction. Soxhlet extraction used by Alam *et al.* may have converted flavonoids into simpler phenolics as we had a much greater content of flavonoids than phenolics while the order in findings of Alam *et al.* (2011) is exactly the opposite.

### Lipid peroxidation inhibition

This assay measures the ability of a sample to inhibit peroxidation that may occur in lipids due to free radicals. Peroxidation of linoleic acid produces peroxy free radicals, which are used to oxidize ferrous ions into ferric ions. The ferric ions so produced form coloured complex, ferric thiocyanate, with thiocyanate ions present in the medium. The formation of the complex is detected spectrophotometrically at 500 nm. Antioxidants will slow down the peroxidation in lipids resulting in low production of peroxides. As a result, there will be less oxidation of ferrous into ferric and therefore less amount of the complex will be formed. Thus, the stronger the antioxidant, the less is the formation of ferric thiocyanate and lower will be the absorbance. In our study, antioxidant activity of the plant samples was plotted as a function of time and compared with a standard antioxidant, BHA (butylated hydroxyanisole) and the results are displayed in Figure 4. The methanolic extract and the fractions in various solvents all showed remarkable ability to inhibit peroxidation in linoleic acid. The plant samples proved to be better antioxidants than BHA. Moreover, the samples remained almost equally effective through the whole duration of the experiment and at the end of the fourth day there was a slight increase in the activity. *Oxalis corniculata*, thus, contains active chemical substances having sustainable antioxidant activities. The CHCl<sub>3</sub> fraction showed an interesting behavior; its activity first decreases reaching its lowest value after 48 hours, and then starts to increase ultimately acquiring a steady state after day 3.

### Conclusion

*Oxalis corniculata* is a common weed distributed throughout Pakistan. In ethnomedicine, it is used as cheap and easily accessible remedy to cure a number of diseases. The plant has good quantities of phenolics and flavonoids, the compounds known for their antioxidant properties. The methanolic extract and its sub-fractions in various solvents, in general, showed good antioxidant activities. Ethyl acetate fraction was found to be most active. Although the plant has already been investigated for its phytochemicals, there is a need for more extensive studies. Assay-guided isolation of natural products from the plant may result into the discovery of new and cheaper therapeutic agents.

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