



ADSORPTION BEHAVIOR AND CORROSION RATE MODEL OF SODIUM CARBOXYMETHYL CELLULOSE (NA-CMC) POLYMER ON ALUMINIUM IN HCL SOLUTION

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ABSTRACT

Adsorption behaviour and corrosion rate of Sodium carboxymethyl cellulose on aluminium was studied. Different concentration levels of Sodium carboxymethyl cellulose (Na-CMC) solution were applied on several prepared and polished aluminium coupons for a corrosion experiment. These coupons were totally immersed in 100ml of 0.5M solution of Hydrochloric acid (HCl) in an open beaker placed in a water bath at varying temperatures for 2-8hrs respectively. At every specified temperature and time interval, the immersed coupon specimen was withdrawn from the test solution, washed, dried and reweighed. The weight loss being the difference in weight of the specimen before and after immersion in the water bath was recorded for every coupon sample. The effects of concentration, temperature and time on corrosion rate were studied and from the various plots, it was observed that increase in the concentration of the inhibitor decreases the corrosion rate. The study of the adsorption behaviour showed that Na-CMC was physically adsorbed on the aluminium coupons and obeyed Freundlich adsorption isotherm with an activation energy of 32K.83j/mol and heat of adsorption of -18.21Kj/mol. The study also established a relationship between the corrosion rate, CR; concentration, C; temperature, T; and time t, through a mathematical model: $CR = 3.8^{-5} * T^{1.1617} - 0.00052t^{0.6176} - 0.0013C^{0.8012}$, the proportion of variance explained (R^2) = 0.8658 (86.58%), From the Results of the experiment and the model, Na-CMC was found to be an active corrosion inhibitor of Aluminium in acidic environment.

Keywords: Adsorption, Aluminium, concentration, Corrosion Rate, Weight Loss

1. INTRODUCTION

Aluminium significantly plays a major role in engineering applications. Most engineering applications involve the use of one form of this metal or the other. Aluminium can be used in different applications to replace other materials like copper, steel, zinc, tin, stainless steel, titanium, wood, paper, concrete and composites [1]. This metal is used in the construction, automobile, in packaging, transport and marine industries. Other applications of aluminium include ladders, high pressure gas cylinders, sporting, furniture, road barriers and signs. Aluminium also

plays major role in areas where light weight application is required like in the construction of towers needed to support the overhead cables for long distant electricity transmission, in aircraft constructions, missile bodies and satellite components etc [2].

Aluminium is applied in acidic, alkaline, neutral and rain water environments. These environments as mentioned by several researchers tend to react with the metal, changing its stable state and this state change causes failure. Failure of this engineering material causes devastating damage materially and in

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some cases loss of life. One of the major causes of these failures include corrosion which is the deterioration of materials by chemical interaction with its environment [3]. The consequences of corrosion effect is so high that scientists, researchers and engineers look for ways of reducing the menacing effect of corrosion. Corrosion control is a process aimed at reducing the corrosion rate of a particular process to a tolerable level. The extent of control depends on the type and method of control used [4, 5, 6]. Corrosion can be controlled by the use of inhibitors, coatings and linings, material selection and cathodic protection.

Corrosion inhibitors are chemicals that are added to the controlled environment to reduce corrosivity of the environment. These inhibitors react with a metallic surface or the environment they are exposed to, giving the surface a certain level of protection [7, 8]. Some of these polymeric inhibitors are biodegradable, cost effective, affordable and soluble in most media used. These can also be produced in large quantities for industrial and commercial use. Because of the hazardous effects of inorganic inhibitors and their health implications, there is the need for an organic material which is environmentally friendly, soluble in most media and at the same time cheap, affordable and readily available for use as protective soluble substances in coating or protecting these engineering materials from collapse [9]. These features made Sodium Carboxymethyl cellulose to be considered as good corrosion inhibitor especially when large scale protection is needed. Several authors like Obot, Solomon, Abdalla, Okafor, Umoren, Nwanonyi etc. had investigated corrosion inhibition of aluminium using several inhibitors and under different environments [7, 8, 10, 11, 13, 25] but none had been able to establish a theoretical model between the corrosion rate and the independent variables like concentration, time and temperature. In this work, we investigated the corrosion inhibition of aluminium using Na-CMC and establish a relationship between corrosion rate, time temperature and inhibitor concentration through a theoretical model and the model results corroborated with that of the experiment.

2. MATERIALS AND METHODS

2.1 Materials and Equipment:

Materials used include AA1060 type of aluminium sourced from Alo Aluminium Ltd, Sodium carboxymethyl cellulose (produced by Qualikem Fine

Chem Pvt Ltd, Deihi, India and purchased from Lavans Chemicals Ltd, the company's distributor in Nigeria, emery papers, water bath, ethanol, acetone, distilled water, HCl (all analytical grades, sourced from GeoChem Ltd), desiccators, beakers, conical flask, timer, glass hook, and digital weighing balance.

2.2 Metal Specimen Preparation

Aluminium of 98% purity and AA1060 type was used and was sourced from Alo aluminium Ltd. Aluminium specimens were mechanically press-cut into coupons of dimensions 5x3x1mm. They were degreased by washing them in ethanol and then polished with emery paper. The metal samples were then rewashed in ethanol rinsed in acetone to dry and then stored in a desiccator before use.

2.3 Experimental Method

Weight loss measurement and theoretical modelling were used to determine the corrosion rate model and adsorption behaviour of Na-CMC on aluminium substrate.

2.3.1 Weight Loss Measurement

In this experiment, a prepared aluminium metal coupon was totally immersed in 100ml of 0.5M HCl acid solution in a 120ml capacity glass beaker with the aid of plastic thread and glass rod. The beaker was placed in a water bath equilibrated at 35°C, 45°C, 55°C, and 65°C respectively [10]. At the end of stipulated time intervals (2h, 4h, 6h and 8h respectively), the metal specimen was withdrawn from the test solution, washed, dried and reweighed [11]. The weight loss was determined as the difference in weight of the metal specimen before and after immersion in the test solution respectively. The test was conducted in triplicates to guarantee the reliability of the results and the average value of the weight loss (W_1) was taken. This procedure was repeated with varying concentrations of the inhibited solution. The following parameters determined from the results obtained are stated below;

Weight loss (W_1): This is regarded as the material loss (reduction in size, shape and weight) due to degradation caused by corrosive agent. Mathematically, it can be evaluated according to Equation (1):

$$\text{Weight Loss } (W_1) = M_1 - M_2 \quad (1)$$

where, M_1 = average weight in grams before immersion, M_2 = average weight in grams after immersion.

2.3.2 Corrosion Rate (R)

This is corrosion parameter that gives physical and visual information regarding the extent of penetration of the corrodent into the metal surface damage with respect to variation in time (3d). Usually, it is expressed in terms of loss of material thickness per unit time and engineers express it in microns per year ($\mu\text{m}/\text{y}$) or mils per year (mpy) or inches per year (ipy) millimetre per year (mm/yr) or milligrams/square decimetre or mm per year. Mathematically, it is expressed according Equation (2) stated below:

$$\text{Corrosion Rate (R)} = \left(\frac{87600\Delta W}{\rho A t} \right) \quad (2)$$

where, ΔW = weight loss in gram, ρ = density of the metal sample (g/cm^3), A = surface area of exposed metal sample (cm^2) and t = time of exposure (in hrs) [12].

2.3.3 Degree of Surface Coverage (θ)

This corrosion parameter measures the extent or amount of corrosion site on the metal surface covered by the inhibitor. It is expressed using equation (3):

$$\theta = (R_b - R_i) / R_b \quad (3)$$

where, R_b = corrosion rates of blank solution, R_i = corrosion rates of inhibited solution.

2.3.4 Adsorption Isotherm

This is a parameter used to evaluate the mode of adsorption inhibition process. All the adsorptions isotherms were tested and plotted but only Freundlich adsorption isotherms has the highest R^2 value and was used in this study. This isotherm is expressed as:

$$\text{Freundlich Isotherm } \log \theta = \log K_a + n \log C \quad (4)$$

2.4 Theoretical Modelling

A mathematical model was designed during the corrosion inhibition investigation of the effects of Sodium Carboxymethyl Cellulose on aluminium coupons under acidic conditions using Phillip Sherrod's NLREG version 6.3 software. This model shows the relationship between Temperature **T**, Time **t**, concentration **C** and corrosion rate **CR**, and was derived under the following conditions.

Total no of observations the experiment = 96

No observations used in the NLREG = 90

No of observations used to test the mathematical model = 6 was used to validate the model.

The model has the following functions:

$$\text{Corrosion Rate (CR)} = K_1 T^{a_1} + K_2 T^{a_2} + K_3 T^{a_3} \quad (5)$$

Where k_1, k_2, k_3, a_1, a_2 and a_3 are the parameters.

3. RESULTS AND DISCUSSION

3.1 Weight Loss (Gravimetric) Measurements

The dissolution and inhibition of aluminium coupon in 0.5M HCl acid solution without and with different concentrations of Na-CMC was investigated at varying temperatures (35°C , 45°C , 55°C , and 65°C) and time respectively using weight loss measurement technique. The results obtained were presented in Table 1. It can be seen from Table 1 that dissolution of aluminium in the test solution increased with time and temperature but the increment was much recorded in the blank specimen solution compared with the inhibited solutions [13]. It was observed that the relationship between the material loss and time at different inhibitor concentration and temperatures respectively is linear. In addition, it was seen that introduction of Na-CMC in the blank solution reduced the corrosive nature of the system thereby regulating corrosion damage on the aluminium surface. Hence, this regulatory action of aluminium corrosion in HCl acid solution by Na-CMC was concentration dependent [14]. Critical evaluation of Table 1 revealed the influence of temperature on the corrosion behaviour of aluminium and inhibitive performance of Na-CMC on aluminium exposed in 0.5M HCl acid solution [15]. It was observed that increment in temperature boosted the damaging action of the corrosive agent (chloride ions) on the aluminium surface and even thwarted the protective effect enjoyed by Na-CMC at low temperature. The results obtained suggest that Na-CMC possess inhibitive property and can serve as inhibitor for aluminium in HCl environment that requires low temperature exposure.

3.2 Corrosion Rate

The corrosion rate is closely related to weight loss in the sense that increasing the concentration of sodium carboxymethyl cellulose decreases the corrosion rate. Calculated values of corrosion rate against inhibitor concentration (Na-CMC) is shown in Table 2 and Figures 3 and 4. From Figures 3 and 4, corrosion rate is highest on the blank coupons when there was no sodium carboxymethyl cellulose added to the aluminium coupon specimen.

Table 1 Calculated values of weight loss for Aluminium corrosion in 0.5M HCl in the absence and presence of Na-CMC at different time and temperatures.

Time (h)	Conc (g/L)	Weight loss (g)			
		35°C	45 °C	55 °C	65 °C
2	Blank	0.050	0.085	0.131	0.156
	0.200	0.025	0.058	0.099	0.142
	0.400	0.020	0.048	0.092	0.135
	0.600	0.017	0.042	0.086	0.124
	0.800	0.014	0.035	0.073	0.104
	1.000	0.007	0.020	0.051	0.073
4	Blank	0.071	0.107	0.146	0.229
	0.200	0.038	0.066	0.112	0.194
	0.400	0.032	0.060	0.102	0.188
	0.600	0.028	0.055	0.098	0.173
	0.800	0.025	0.048	0.084	0.152
	1.000	0.015	0.035	0.063	0.115
6	Blank	0.081	0.125	0.168	0.310
	0.200	0.050	0.079	0.129	0.270
	0.400	0.043	0.072	0.125	0.258
	0.600	0.039	0.069	0.101	0.240
	0.800	0.035	0.061	0.085	0.225
	1.000	0.021	0.040	0.072	0.175
8	Blank	0.090	0.140	0.180	0.394
	0.200	0.065	0.098	0.134	0.345
	0.400	0.058	0.091	0.125	0.314
	0.600	0.053	0.082	0.115	0.285
	0.800	0.048	0.075	0.107	0.270
	1.000	0.030	0.052	0.085	0.225

The corrosion rate gradually reduces as the concentration of the inhibitor (Na-CMC) is increased and it is lowest when the concentration of 1g/L solution of the inhibitor is applied. This is as a result of the increase in the coverage of the surface area of the aluminium coupons as the concentration of the inhibitor increases. The coverage would not allow the harsh acidic environment to attach aluminium sample by so doing reducing the effects and potency of the acid on attacking the metal. Sodium Carboxymethyl Cellulose affects both the weight loss, corrosion rate and inhibition efficiency in corrosion experiment.

3.3 Temperature Effect on Corrosion of Aluminium

Temperature increases the rate of mobility of molecules in corrosion experiment. As the temperature increases, the rate of disorder of the molecules also increases [16]. This disorder weakens the bonding of the inhibitor on the aluminium surface thereby exposing the coupons to attack by the acid [17]. Temperature changes in corrosion experiment play a

vital role in the determination of the corrosion rate. At a particular concentration and fixed time interval, increase in the temperature of the environment increases the rate of corrosion of the aluminium coupons. From Table 2, it is evident that at 2h time interval when the inhibitor concentration is 0.2g/L, corrosion rate at the temperature of 35°C is 27.04mm/yr.

The rate increased to 62.73mm/yr as the temperature reached 45°C and to 153.57mm/yr at 65°C. Figures 5 and 6 show the relationship between corrosion rate and temperature and from the figures shown, it is observed that corrosion rate increased as the temperature increased.

This increase in temperatures breaks down the Van der Waal forces holding the molecules together with the aluminium interface and allow the Na-CMC to dissolve into solution.

This dissolution reduces the level of surface coverage of the coupon by the inhibitor and at the same time increasing the strength of acidic medium to attack the aluminium specimen coupon. This dissolution exposes

the aluminium surfaces for attack by the acidic environment (HCl). The rate of attack is more prominent as the temperature is increased and highest on the blank coupons as can be seen from the plot. Therefore, increase in the temperature of the environment increases the corrosion rate and vice versa. Increase in the temperature increases the corrosion rate as seen from the model equation and decreases as the temperature is lowered [18]. Therefore the corrosion rate of the model and that of the experiment follow the same trend in temperature relationship.

3.4 Adsorption Studies on Corrosion of Aluminium

Adsorption isotherm was used to evaluate the adsorption process due to the fact that adsorption is a surface phenomenon. Adsorption Isotherm is a graph that expresses the variation in the amount of molecules adsorbed by the adsorbent with concentration at constant temperature [19].

In order to determine which adsorption isotherm to adopt in the work, we use the values of the calculated surface coverage (θ) to fit in the already existing isotherms and the one with the highest R^2 values were used to determine the thermodynamic parameters governing the inhibitor adsorption. Freundlich Isotherm has the highest R^2 value and therefore fits in our experiment.

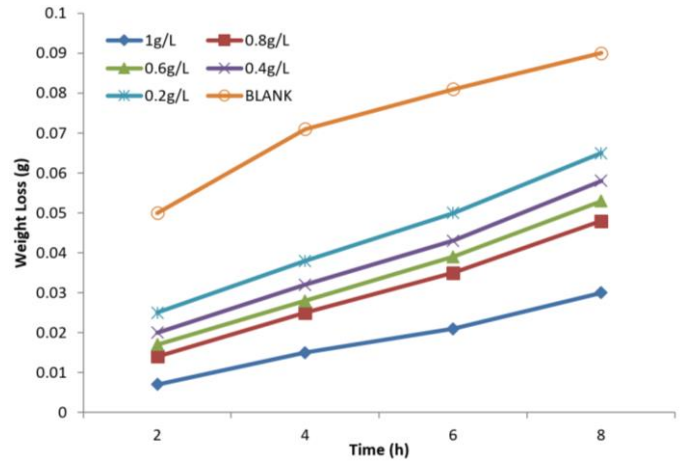


Fig.1: Variation of weight loss against time at varying concentration at 35°C

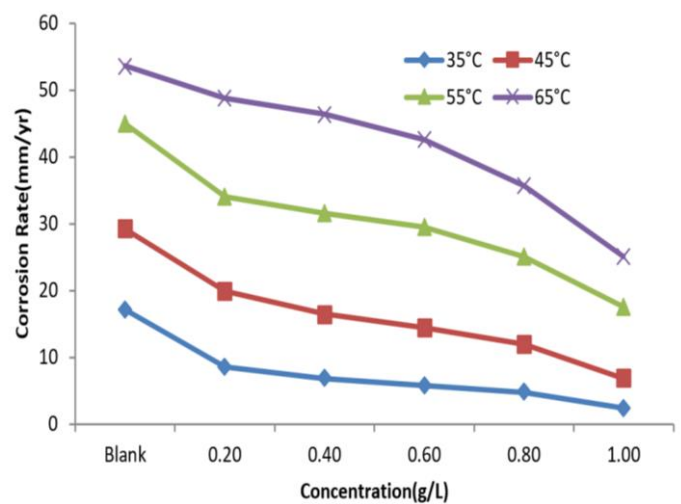


Fig 3: Variation of corrosion rate against concentration at varying temperatures for 2h

Table 2: Values of concentration against corrosion rate and Surface Coverage at different temperature and Time

Time(hrs)	Conc(g/L)	Corrosion rate(mm/yr)				Surface Coverage			
		35°C	45°C	55°C	65°C	35°C	45°C	55°C	65°C
2	Blank	17.19	29.22	45.03	53.62	-	-	-	-
	0.2	8.59	19.94	34.03	48.81	0.50	0.32	0.24	0.09
	0.4	6.87	16.50	31.62	46.41	0.60	0.44	0.30	0.13
	0.6	5.84	14.44	29.56	42.62	0.66	0.51	0.34	0.21
	0.8	4.81	12.03	25.09	35.75	0.72	0.59	0.44	0.33
	1.0	2.41	6.87	17.53	25.09	0.86	0.76	0.61	0.53
4	Blank	12.20	18.39	25.09	39.36	-	-	-	-
	0.20	6.53	11.34	19.25	33.34	0.46	0.38	0.23	0.15
	0.40	5.50	10.31	17.53	32.31	0.55	0.44	0.30	0.18
	0.60	4.81	9.45	16.84	29.73	0.61	0.49	0.33	0.24
	0.80	4.30	8.25	14.44	26.12	0.65	0.55	0.42	0.34
	1.00	2.58	6.02	10.83	19.77	0.79	0.67	0.57	0.50
6	Blank	9.28	14.32	19.25	35.52	-	-	-	-
	0.20	5.73	9.05	14.78	30.94	0.38	0.37	0.23	0.13
	0.40	4.93	8.25	14.32	29.56	0.47	0.42	0.27	0.17

Time(hrs)	Conc(g/L)	Corrosion rate(mm/yr)				Surface Coverage			
8	0.60	4.47	7.91	11.57	27.50	0.52	0.45	0.40	0.23
	0.80	4.01	6.99	9.74	25.78	0.57	0.51	0.49	0.27
	1.00	2.41	4.58	8.25	20.05	0.74	0.68	0.57	0.44
	Blank	7.73	12.03	15.47	33.86	-	-	-	-
	0.20	5.59	8.42	11.52	29.65	0.32	0.30	0.25	0.12
	0.40	4.98	7.82	10.74	26.98	0.36	0.35	0.31	0.20
	0.60	4.55	7.05	9.88	24.49	0.41	0.42	0.36	0.27
	0.80	4.12	6.45	9.20	23.20	0.47	0.46	0.41	0.31
	1.00	2.58	4.47	7.30	19.34	0.67	0.63	0.53	0.43

Table 3 shows the values of inhibitor concentration(g/L), temperature and degree of surface coverage used to find the isotherm the experiment followed.

Freundlich Isotherm $\log\theta = \log K_a + n\log C$ (4)

Where K_a is the equilibrium constant of the adsorption process, C the inhibitor concentration, θ is the surface coverage. It is a known fact that K_a is the strength between adsorbate and adsorbent [20]. The best fitted straight line was obtained from the plot of $\log\theta$ against $\log C$, that is the Freundlich Isotherm with an R^2 value of 99.88% at temperature 65°C . Therefore, this corrosion test we carried out obeyed Freundlich adsorption isotherm.

Freundlich Adsorption Isotherm

The Freundlich equation or Freundlich adsorption isotherm is an empirical relation between the concentration of a solute on the surface of an adsorbent to the concentration of the solute in the inhibitor with which it is in contact [21]. This shows a plot of the log of the surface coverage (Log θ) against the log of concentration (Log C). Table 4 shows the values of Log C and the corresponding values of Log θ at different temperatures. The plot of log θ against log C gives the slope as n and the intercept as log K . The R^2 values of the plot of Log θ against Log C at various temperatures are shown in Fig. 7. The highest R^2 of 99.85% was recorded at a temperature of 65°C.

3.5 Thermodynamic Considerations of Corrosion Aluminium

3.5.1 Activation Energy

In corrosion studies, the corrosion rate is related to temperature by the Arrhenius equation 6. This is the minimum amount of energy required to break the van de Waal forces holding the molecules of the inhibitor out of the surface of the aluminium before corrosion can take place [22].

$$\text{Log} \frac{R_b}{R_i} = \frac{E_a}{2.303R} \left[\frac{1}{T_1} - \frac{1}{T_2} \right] \quad (6)$$

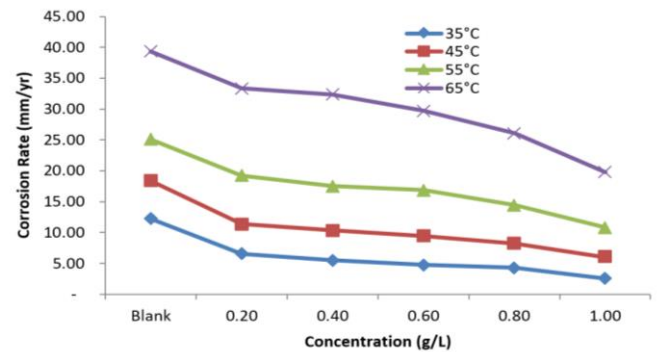


Fig 4 Variation of corrosion rate against concentration at varying temperatures for 4h

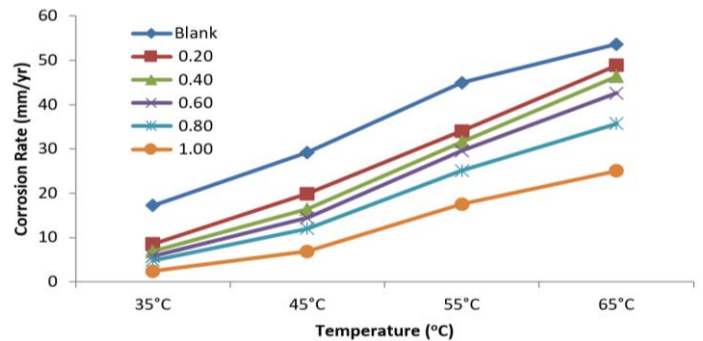


Fig 5 Variation of corrosion rate against temperature at varying concentration for 2h.

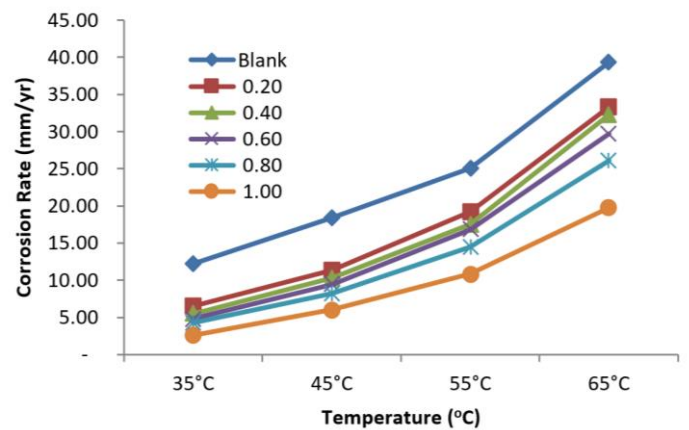


Fig. 6 Variation of corrosion rate against temperature at varying concentration for 4h.

Table 3 Inhibitor concentration and calculated surface coverage

concentration(g/L)	Surface Coverage			
	35°C	45°C	55°C	65°C
0.0000	0.0000	0.0000	0.0000	0.0000
0.2000	0.5000	0.3176	0.2443	0.0898
0.4000	0.6000	0.4353	0.2977	0.1346
0.6000	0.6600	0.5059	0.3435	0.2051
0.8000	0.7200	0.5882	0.4427	0.3333
1.0000	0.8600	0.7647	0.6107	0.5321

Table 4 Calculated values of log C vs log θ

Log C(g/L)	Log θ			
	35°C	45°C	55°C	65°C
-0.699	-0.301	-0.4981	-0.6121	-1.0467
-0.3979	-0.2218	-0.3612	-0.5262	-0.871
-0.2218	-0.1805	-0.2959	-0.4641	-0.688
-0.0969	-0.1427	-0.2305	-0.3539	-0.4772
0	-0.0655	-0.1165	-0.2142	-0.274

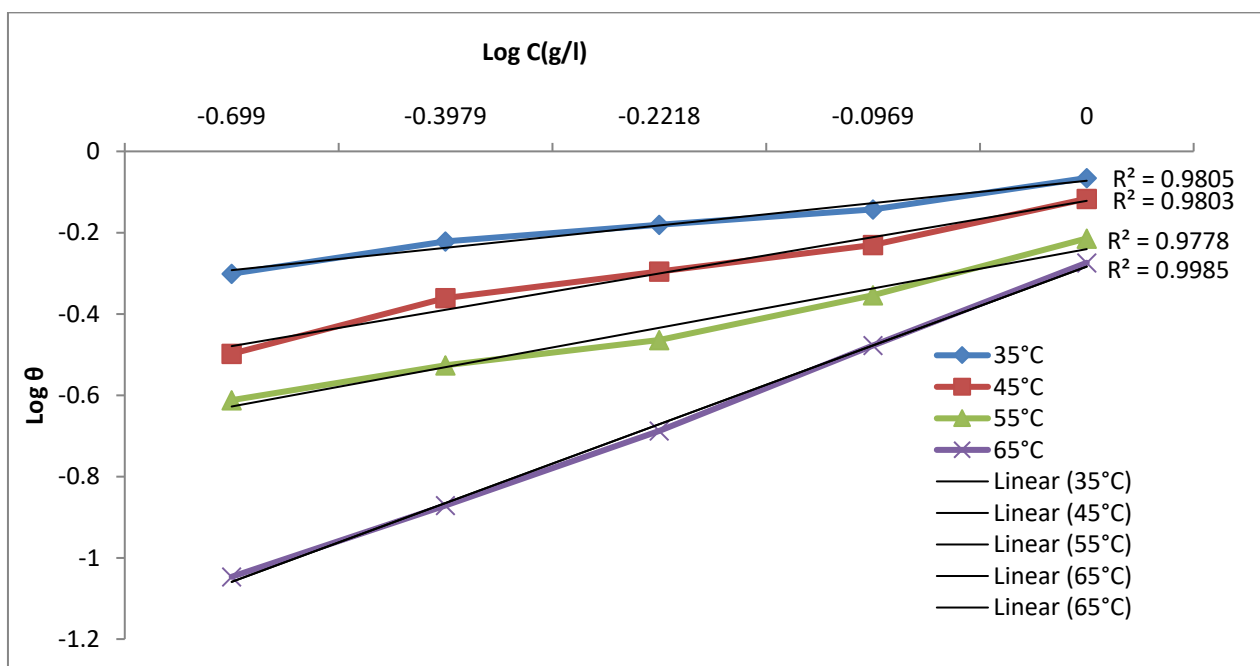


Fig. 7 Freundlich adsorption isotherm plot of log θ versus log C for sodium carboxymethyl cellulose at different temperatures.

In Equation (6) R_b is the corrosion rate of the blank coupon and R_i is the corrosion rate of the inhibitor determined from the weight loss measurement, E_a is the apparent activation energy, R the molar gas constant ($8.314 \text{ J K}^{-1} \text{ mol}^{-1}$) and T_1 and T_2 is the absolute temperatures of the blank coupon and the inhibitor respectively [23]. The calculated values of the apparent activation energy, E_a is shown in Table 5 below.

From Table 5, the calculated values of E_a in the presence of different concentrations of the inhibitor

were higher than that in its absence. This is because of the formation of an adsorptive film of a physical electrostatic character of sodium carboxymethyl cellulose and this adsorbed Na-CMC created a physical barrier against attack of the aluminium by the acidic medium. The activation energy increases as the inhibitor concentration increases. This is because of increase of adsorptive films on the surface of the aluminium coupons as the concentration increases [24].

Table 5: Calculated values of activation energy (E_a) and heat of adsorption (Q_{ads}) for aluminium in 0.5M HCl for the different concentrations of Na-CMC for 2hrs.

Concentration (g/L)	Activation energy (E_a) (kJ/mol) 35°C-65°C	Heat of adsorption (Q_{ads}) (kJ/mol) 35°C-65°C
Blank	32.83	
0.2	50.13	-18.21
0.4	55.12	-17.06
0.6	57.35	-16.58
0.8	57.88	-14.05
1.0	67.60	-12.37

From theory, if the activation energy of blank coupon (E_{ab}) is less than the activation energy of the inhibitor (E_{ai}), then, the mechanism is physisorption. From Table 5 below, the adsorption of Na-CMC is typical physisorption because E_{ab} is less than E_{ai} .

3.5.2 Heat of Adsorption:

Heat of adsorption is the heat evolved when a given amount of the inhibitor is adsorbed on the surface of the aluminium coupon. An estimate of heat of adsorption was obtained from the trend of surface coverage with temperature as shown in equation 7 [25].

$$Q_{ads} = 2.303R \left[\log \left(\frac{\theta_2}{1-\theta_2} \right) - \log \left(\frac{\theta_1}{1-\theta_1} \right) \right] \times \frac{T_1 T_2}{T_2 - T_1} \quad (7)$$

The calculated value of the heat of adsorption of Na-CMC on the aluminium coupon is shown in table 5. More importantly, the heat of adsorption is used to confirm if the reaction is chemisorption or physisorption. It is believed that the mechanism is physisorption if heat of adsorption is less negative than -40kJmol^{-1} and chemisorption, if enthalpy of adsorption is more negative than -100kJmol^{-1} [26]. The heat of adsorption in this our experiment is less than -40kJmol^{-1} , therefore the adsorption of sodium carboxymethyl cellulose on the aluminium coupon is typical physisorption.

3.6 Corrosion Rate Model

Corrosion modelling expresses the relationship between the independent variables; concentration, temperature and time and the dependent variable corrosion rate in a mathematical format. The model equation shows the relationship between corrosion rate, concentration, temperature and time interval of the experiment. R^2 values and Prob(t) are statistical values used to validate the overall significance of the model produced. Also, the data from the experiment were substituted in the model and were used to validate the model equation. The mathematical model equation developed in this our work relates the

Corrosion Rate (CR) with the variables used in the NLREG model. The independent variables are temperature (T), Concentration (C) of the inhibitor used, time (t) intervals used in the corrosion experiment while the dependent variable is Corrosion Rate (CR). The relationship between the variables was presented with a model equation as given in equation 5.

$$\text{Corrosion rate (CR)} = K_1 T^{a_1} + K_2 T^{a_2} + K_3 T^{a_3} \quad (5)$$

Where k_1, k_2, k_3, a_1, a_2 and a_3 are parameters.

The result of the parameters/constants obtained from the plot of the variable functions using the Non-linear regression analysis software are presented in Table 6 Substituting the values of the parameters of Table 6 in equation (5), gave equation

$$CR = 3.8525^{-5} T^{1.1617} - 0.00052 t^{0.6176} - 0.0013 C^{0.8012} \quad (8)$$

Therefore, equation (8) is the theoretical model derived from the experiment and this has a proportion of variance explained (R^2) = 0.8658 (86.58%). This R-squared (R^2) value measures the strength of the relationship between the model and the dependent variable on a convenient 0 – 100% scale. If the function perfectly predicts the observed data, the value of this statistic will be 1.00 (100%). If the function does no better a job of predicting the dependent variable than using the mean, the value will be 0.00. Our model gave an R^2 value of 86.58% which significantly showed that the strength of the relationship between the experimental variables; Corrosion Rate (R), Temperature (T), Time (t), and Concentration (C) as computed in the NLReg model software is very high.

Table 7 and Fig. 8 show the comparison between the experimental and model data. It is observed that increase in temperature increases corrosion rate for both the experiment and model. Therefore the model confirms the results of experimental work. In fig. 8, the two graphs seem to align together.

Table 7 Comparison of the graph of Corrosion Rate from both the Model and the Experiment.

Temp(°C)	Time(hrs)	Conc(g/L)	Corrosion Rate (mm/yr)	
			Model	Experiment
35	2	0.8	0.000549999	0.000466667
35	6	1	0.000425211	0.000233333
45	4	0.4	0.00138866	0.001
45	8	0.2	0.000997357	0.000816667
55	2	0.6	0.00242188	0.002866667
55	6	0.4	0.00214319	0.001388889
65	4	0.2	0.00335657	0.003233333
65	8	0.8	0.002000024	0.00225

The slight deviation may be caused as a result of environmental factors which affect only the experiment as it is being conducted. The environmental factors include: Human error of omission and commission as readings were being taken, power and other factors which affect the experiment and this brought about the slight deviation from the model as can be seen in Fig 8

Table 6 Calculated value of Corrosion rate parameters used from the model.

Parameter	Final estimate
k ₁	3.85250959E-005
k ₂	-0.00051603286
k ₃	-0.00126085193
a ₁	1.16172466
a ₂	0.617626634
a ₃	0.801175129

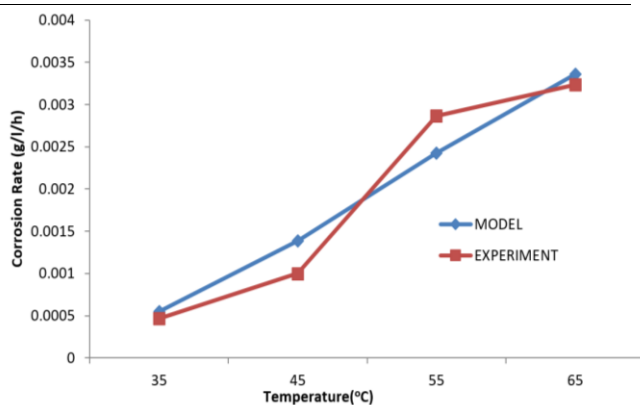


Fig. 8: Variation of Corrosion Rate Against Temperature from the model & that of the experiment.

4. CONCLUSION

Sodium Carboxymethyl cellulose is an active corrosion inhibitor for Aluminium on acidic environment. Increasing the concentration of the inhibitor decreases the corrosion rate and vice versa. Also, increasing the temperature of the operating environment increases the corrosion rate.

The relationship between the concentration of the inhibitor, the temperature of the operating environment, the exposure time and the corrosion rate is represented with a model:

$$CR = 3.8525^{-5}T^{1.1617} - 0.00052t^{0.6176} - 0.0013C^{0.8012} \tag{9}$$

the proportion of variance explained (R²) = 0.8658 (86.58%) and the prob(t) is 0.000001.

The results of thermodynamic parameters indicated that the adsorption of Na-CMC on the aluminium surface obeyed Freundlich adsorption isotherm and was a typical of physisorption.

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