Toward zero waste production in the paint industry

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Abstract

Wastewater is generated in the batch production process of water-based paints when vessels and filling lines are washed between batches. This results in a diluted paint wash water stream. The wash water is currently treated in a flocculation process using ferric chloride as a coagulant.

An opportunity was identified for re-using wash water from white, non-textured paints. It is necessary to stabilise the wash water with biocides. This water is then used in place of fresh feed water for lower quality paints. Since the wash water originates from high quality products, adding this water to the lower quality paints results in a quality improvement. In addition, treatment costs of the wash water are reduced. This wash water represents approximately 65% of total effluent from the plant.

With respect to the remainder of the wash water, the concentration of paint solids in the wastewater stream can vary widely. For effective treatment different solids content in the wastewater requires different dosage levels of the coagulant. If ferric chloride is used as the coagulant overdosing results in waste that must be disposed of as toxic. Ferric chloride dosing can be more easily controlled by using a dilute feed solution. This paper presents the findings of the effect of concentration of the coagulant on the coagulation process. Both the solids content of the wash water and the concentration of the coagulant have been found to be important variables to ensure efficient flocculation and coagulant utilisation.

Introduction

Although the type of waste generated by different industries varies dramatically the principles in waste minimisation are generic.

An approach known as the hierarchy of waste is being applied. The basic principle is that to avoid waste generation is far easier than to reprocess it later (Cano-Ruiz and McRae, 1998). The next level in the hierarchy is to simply reuse the waste material with almost no processing. A common example of this is the refilling of bottles after simply washing them. Recycling is the next level down and it implies that the waste stream is reprocessed to make it useful again. Again the analogy of a glass bottle being melted down to be formed into a new shape is useful. Potentially the most difficult waste reduction strategy is to reprocess the waste in such a way that it can be used in another unrelated process. This could however also prove to be a simple solution. It is clear that the further from the source of the waste one tackles the problem of waste minimisation, the more energy, technology and cost is required, and hence the more difficult it becomes to strive toward a zero waste system. The major driver for waste reduction is cost. The diagram that follows is a representation of the hierarchy of waste.

In the drive for zero waste Barloworld Plascon SA (Pty) Ltd has undertaken a number of initiatives to eliminate waste at source.

Two of these will be discussed in this paper, namely recycling wash water and the way in which effluent is treated prior to disposal.

Reusing the wash water is a desirable way to minimise the waste generated. This could be done in two different ways: the water could be reused as wash water, however this would result in a fairly fast deterioration of the cleaning of vessels and lines, as the paint is a fairly concentrated suspension. The other alternative is to explore re-using the wash water instead of fresh water feed. The major concern that needs to be addressed is product quality which must not be affected by the re-use of the wash water.

For example depending on the origin of the wash water it may contain solids with undesirable properties like texture or colour which could affect product quality. Wash water with texture or colour could only be used in products with texture and colour. Moreover the wash water ferments and storage space is limited. In the light of this and the fact that the paint production schedule runs according to market demand for different products, not all of the water can be reused. The remainder must continue to be disposed of once it has been treated.

The wastewater generated is comprised of water and paint constituents (mainly solids that are suspended in the wastewater). This wastewater cannot be sent directly into the public sewage system because it does not meet the effluent discharge standards.
such as turbidity and solids content.

Solids are removed from the wastewater by means of coagulation and flocculation using ferric chloride. The solids content of the water is variable and the amount of ferric chloride required to flocculate the water is proportional to the solids content of the water. This sometimes leads to ferric chloride over-dosing. When this occurs the water must be disposed of as toxic due to the high Fe³⁺ concentration. It is expensive to dispose of toxic waste, so that eliminating overdosing represents a considerable cost saving to the company. The sludge must also be disposed of. The thinking around the re-using or recycling of the sludge waste into a product that is useful to the paint or other industries is still in its infancy. The solids in paint can be divided into two categories: particles that settle out under gravity in a matter of hours (ca. 1 mm in diameter) and particles that form a colloidal suspension (ca. 0.1 mm in diameter) (Peavy, Rowe and Tchobanoglous, 1985). Coagulation is defined as the destabilisation of a colloidal suspension, and flocculation as the subsequent agglomeration of small particles into flocs that settle out relatively fast. These two processes are combined to separate colloidal suspensions into a sludge that contains the solids and clear liquid.

Coagulation is achieved by adding a coagulant to the water to be treated. Different types of coagulants include organic surfactant type molecules or inorganic salts, usually of iron or aluminium. When added to water these salts hydrolyse to form a range of type molecules or inorganic salts, usually of iron or aluminium.

When added to water the ferric ion (Fe³⁺) in ferric chloride (FeCl₃.6H₂O) hydrolyses to form ferric hydroxide, see Eqs. 7 and 8.

\[
\begin{align*}
\text{Fe}^3+ + 3\text{OH}^- & \leftrightarrow \text{Fe(OH)}_3 (\text{am}) \quad \text{log} \ K (25^\circ\text{C}) = -6 \quad [3] \\
\alpha- \text{FeOOH} (c) + \text{H}_2\text{O} & \leftrightarrow \text{Fe}^{4+} + 3\text{OH}^- \quad \text{log} \ K (25^\circ\text{C}) = -37.1 \quad [8]
\end{align*}
\]

In order to obtain comparable results a solid content of 80 g of solid per 1 000 g of wastewater was selected because it was the most dilute sample encountered. The wastewater generated in order to obtain comparable results. A solid content of 80 g of solid per 1 000 g of wastewater, which contains six stirrers with paddles according to the general procedure. (Faust and Aly, 1983; Hammer, 1977; Kemmer, 1988). Six beakers are used in an experiment, each containing 400 mL of wastewater. 20, 40, 60, 80, 100 or 120 mL of flocculent solution is added to the beaker. On addition of coagulant solution the wastewater is agitated at a speed of 250 rpm for one minute to aid quick dispersion of the flocculent into all parts of the wastewater in the beaker. Formation of flocs is observed after about one and half minutes. The speed is reduced after ten minutes and six minutes later it is finally stopped so that sedimentation of flocs can occur. The contents of the beaker are allowed to settle for a further thirty minutes, so that the water above the sludge can be recovered.

Determination of the solid content of wastewater was done by drying a known mass of 20 g of wastewater in oven at 130°C for one hour. The solid content of wastewater of different samples varies and it was standardised to 80 g of solid per 1 000 g of wastewater, in order to obtain comparable results. A solid content of 80 g of solid per 1 000 g of wastewater was selected because it was the most dilute sample encountered. The wastewater generated in the paint production process is collected and stored in a vessel; it is agitated continuously for about two hours to ensure even dispersal of solid in colloids before sampling for the purposes of this study. The wastewater sample is kept in a sealed 5L container, and stirred at 300 rpm for about one hour before the solid content of wastewater is measured.

Metal ion concentrations of water recovered were determined using a Varian Flame Atomic Adsorption spectrometer (Model:
Spectra AA 55B). In the measurement of ferric ions, acetylene is used as fuel with air as support and the flame stoichiometry is oxidizing. For aluminium ions, acetylene is used as fuel the support is nitrous oxide. The flame stoichiometry is reducing.

Turbidity was measured with an infra-red Hanna Portable Microprocessor Turbidity Meter; the unit of measurement is Formazin Turbidity Unit (FTU). The pH meter is a CRISON micro pH 2002 model. Calibration is carried out using technical buffer solutions of 4 and 7.

Results and discussion

Recycling wastewater

The waste from the Barloworld Plascon SA plants is mainly a result of using water to wash the tanks, pipes, and filling machines used to produce and package the products. The products are produced in batches, with different products scheduled in the same vessels. This waste stream is therefore made up of dilute paint which until recently was all treated in a flocculation process with ferric chloride as the coagulant.

An opportunity was identified where the non-textured white wash water could be captured and reused in low quality white paints. As the wash water will be reused in low quality products, washings from low quality products are also captured.

During preliminary attempts to reuse the wash water concerns were raised due to biological and product contamination. Paint contains many organic ingredients, most notably cellulose thickeners and polymers, and is therefore prone to microbiological attack. There is an inherent risk that if the water is contaminated, and then used, the product made from the water would rapidly deteriorate with disastrous consequences on product quality and for the receiving market. In order to minimise risk, the process of collecting the wash water and treating it is almost fully automated and biocide is added to the water. In addition, an outside company tests the wash water collected for contamination every three weeks and a visual check of the water is done before use. The combination of all of these measures ensures that the wash water recovery system functions correctly and that there is no risk to the consumer.

The wash water storage vessel is agitated in order to prevent the heavier solids from settling and blocking the transfer lines. On average a batch of water is collected over about 3 days.

The volumes of feed water that the paint making process requires are dramatically larger than the amount of wash water produced. This provides an opportunity to completely empty the wash water storage vessel every time a batch of paint is made from the wash water. This has two distinct advantages in that the agitator is not required to operate continually, and more importantly, one batch of wash water is not mixed with the next batch. This reduces energy consumption, and further minimises the risk of contamination. The characteristics of the wash water are given in Figs 2 and 3.

The wash water feed contains an average of 12% solids, ranging from 7.6% to 18.6%, see Figure 2. The percentage solids in paint varies from 50 to 60%. These solids are components of the products made so that recycling the wash water back into the process results in an improved product quality. In addition some savings of virgin raw materials are made. In particular, less rheology modifiers are required. The exact reduction varies from batch to batch as the mix of raw materials in the wash water varies.

The composition of the wash water depends on the production schedule, in other words, what is being washed away. Moreover, the exact volume of wash water is variable since the washing is done manually. At present precise information on the quantity and nature of the wash water is not available. If this information were known, further reductions in raw materials would be possible.

The pH of the wash water varies from 8.7 to 9.5 and is shown in Fig. 3. This is somewhat higher than the pH of feed water. Since some of the components of the product are also basic, the high pH of the wash water does not affect the final product quality.

Results of the wastewater reuse to date

The products made from the wash water are carefully monitored over an extended period to determine if there are any long term consequences for the product as a result of the use of the wash water. To date, the results have been excellent. All of the critical parameters that are monitored are well within the required specifications and have remained stable. In Figure 4 results for viscosity, one of the critical parameters of the receiving products, are plotted.

The viscosity increases fairly quickly at first and then levels off at a value just below the upper limit. After 115 days the viscosity was still within this limit.

Treatment of waste in order to minimise coagulant usage and sludge production

In Table 1 the results are shown of the experiment that was used to determine the volume of ferric chloride solution required to effect good flocculation of wastewater without overdosing. The same
The volume of wastewater was added to each beaker, namely 400 ml, but the volume of FeCl₃ was varied. Table 1 shows the concentration that corresponds to the volume of ferric chloride dosed into each jar. The concentrations at which flocculation occurs are marked with asterisks (*). It can be seen that the onset of flocculation occurs for ferric chloride between a concentration of 0.91 and 1.17 g/l.

In order to investigate the effect of ferric chloride feed concentration the following solutions were prepared: 52, 109 and 607 g/l. In a similar way to the experiment for which the results are shown in Table 1, various volumes of these solutions were added to wastewater to test for flocculation. The results are given in Table 2.

The onset of flocculation occurs at the dosages indicated by either asterisks (*) or hashes (#), namely: 1.17, 1.21, 1.37 and 1.21 g/l for the 7, 52, 109 and 607 g/l feed solutions. It is interesting to note that the 1.17 and 1.21 g/l tests resulted in coagulation and flocculation, however the 1.13 g/l test did not. This implies that the lower ferric chloride concentration limit for coagulation is between 1.13 g/l and 1.17 g/l. In other words insufficient coagulant was added to the 1.13 g/l beaker for coagulation to occur.

For all of the higher concentration feed solutions the coagulation occurred immediately on addition of the coagulant and the flocs formed were large, this is indicated by #. This is in contrast to the addition of a 7 g/l feed solution that favours gradual growth of fine floc particles. Moreover the liquid recovered on addition of highly concentrated ferric chloride feed was still milky after flocculation, for all cases. The sludge formed using highly concentrated ferric chloride is coarse in nature and lumpy. The coarse flocs tend to float on the surface of the water rather than settle out.

These results can be explained as follows. It was mentioned in the introduction that there is competition between the precipitation of ferric hydroxides and the interaction of hydrolysed ferric species with the colloidal particles of the suspension. In the case of low ferric feed concentrations (7 g/l), the precipitation reaction is slow, so that the ferric polycations react with colloidal particles. For all of the higher concentration ferric chloride feed solutions the precipitation reaction is accelerated by the higher ferric concentration. As soon as the coagulant comes into contact with hydroxide ions in the wastewater, it precipitates. The flocs formed have a much higher ferric content, so that the same amount of coagulant only partially removes the colloidal particles and the water remains milky. In other words the coagulant is used inefficiently.

The results of the 7g/l experiment are given in more detail in Fig. 5. The following data has been plotted: turbidity (which is shown on the left hand vertical axis), pH and the concentration of iron in the clear solution recovered after flocculation (shown on the right hand axis) as a function of the volume of ferric chloride feed solution added. The turbidity of the solution in which ferric chloride was hydrolysed ferric species with the colloidal particles of the suspension. In the case of low ferric feed concentrations (7 g/l), the precipitation reaction is slow, so that the ferric polycations react with colloidal particles. For all of the higher concentration ferric chloride feed solutions the precipitation reaction is accelerated by the higher ferric concentration. As soon as the coagulant comes into contact with hydroxide ions in the wastewater, it precipitates. The flocs formed have a much higher ferric content, so that the same amount of coagulant only partially removes the colloidal particles and the water remains milky. In other words the coagulant is used inefficiently.

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5ppm and reuse standard is 0.04 ppm (GISCLON et al., 2002). The pH of the water recovered was around neutral and within the effluent
disposal limits.

The amount of ferric chloride that needs to be dosed must be
adjusted according to the solid content of the wastewater. Provided
the ferric chloride is dosed correctly, the waste generated is not
hazardous.

Using a low concentration ferric chloride solution has the
disadvantage that larger volumes of wastewater are generated, and
additional feed water is required to dilute the ferric chloride
solution. There is a trade-off between the additional water required
and the reduction in ferric chloride used.

Conclusions

What is the benefit to the environment and to
Barloworld Plascon SA

On the basis of current operation approximately 190kL of wash
water would be reused annually. (There is some unrealised potential
for increasing this.) This represents an estimated 20% of the
total waste stream from the plant. There are savings in terms of
reduced quantity of sludge to be disposed of, reduced ferric chloride required for sludge treatment and reduced quantities of
rheology modifier.

The calculation is indicated more clearly below:

<table>
<thead>
<tr>
<th>Mass of wash water reused since</th>
<th>186 051 kg</th>
</tr>
</thead>
<tbody>
<tr>
<td>September 2003</td>
<td></td>
</tr>
<tr>
<td>Average solids in the wash water</td>
<td>12%</td>
</tr>
<tr>
<td>Solids content in the sludge</td>
<td>98%</td>
</tr>
<tr>
<td>Mass of sludge is therefore</td>
<td>22 782 kg</td>
</tr>
<tr>
<td>(186 051kg x 0.12 / 0.98)</td>
<td></td>
</tr>
<tr>
<td>Disposal cost</td>
<td>R340/ ton</td>
</tr>
<tr>
<td>Savings on disposal cost</td>
<td>R7 700</td>
</tr>
<tr>
<td>Ferric chloride saving (@4c/l treatment cost)</td>
<td>R7 500</td>
</tr>
<tr>
<td>Rheology modifier saving (Acrysol TT935 @ R12.86/kg) on 48 batches of paint</td>
<td>R58 600</td>
</tr>
<tr>
<td>(48 x 95kg/batch x R12.86)</td>
<td></td>
</tr>
<tr>
<td>Total savings from wash water reuse (on an annual basis)</td>
<td>R73 800</td>
</tr>
</tbody>
</table>

Due to the ineffective use of ferric chloride the waste is
classified as hazardous. The avoidance of this waste is therefore of
great environmental importance, even though the financial benefits
are small.

Based on the results obtained from the flocculation experi-
ments the following conclusions may be drawn. The wastewater
generated contains solid content that varies with every batch,
therefore a constant dosage of coagulant cannot be used to treat
paint wastewater. Analysis of solid content of different samples of wastewater should be carried out before coagulation and flocculation.

Dosage of coagulants for wastewater treatment can be based on the
quantity of solids in wastewater, in order to avoid overdosing. The
flocculation of wastewater at high concentration of floculants leads to inefficient use of the coagulant due to the high rate of
hydroxide precipitation.

Using low concentration ferric chloride solutions has the
disadvantage that larger volumes of wastewater are generated, and
additional feed water is required to dilute the ferric chloride
solution. There is a trade-off between the additional water required
and the reduction in ferric chloride used.

The pH and metal ion content of the water recovered are within
the effluent limits, however the turbidity is not.

The sludge produced in the experiment contains the bulk of the
floculent. Therefore a careful study of the sludge is essential
before the sludge can be recycled.

Barloworld is a company dedicated to good environmental prac-
tice.

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