

Improved sulphate removal rates at increased sulphide concentration in the sulphidogenic bioreactor

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Abstract

The product of the biological sulphate reduction is sulphide. High concentrations of molecular H₂S (g) can be inhibitory for microbial activity, especially at a reactor pH of 6 to 7. This paper focuses on the effect of high sulphide concentrations on the sulphate reduction rates. The results of three investigations operating a continuous reactor, a column reactor and batch-test reactors have shown that increased sulphide concentrations have resulted in improved biological sulphate reduction. In all instances the reactor pH was kept at 7.5 to 8.5. It was shown that when the sulphide concentration was 700 mg/l in a continuously operated reactor, the sulphate reduction rate was 12 gSO₄/l·d. When operating batch-test reactors the results showed that when the sulphide concentration increased, to 1 400 mg/l, the volumetric and specific sulphate reduction rates correspondingly increased to 4.9 gSO₄/l·d and 1.5 gSO₄/gVSS, respectively. Thirdly, operating a tall column reactor using H₂ and CO₂ as the energy source, showed that when the initial sulphide concentration of the feed water was 0, 100 and 268 mg/l, the average biological sulphate removals were 650, 1 275 and 1 475 mg/l, respectively. These obtained results indicated that the addition of sulphide to the feed water to the reactor had a positive effect on sulphate removal. Improved sulphate removal results in increased alkalinity production and in an increased reactor pH, which in turn is favourable for a decrease in the redox potential, when a dominant redox couple, like sulphate: sulphide, is present in a reactor.

Keywords: alkalinity, pH, Redox value, sulphate reduction rate, sulphide toxicity

Introduction

Acid mine drainage (AMD) is the result of mining activities, due to bacterial oxidation and exposure to oxygen and water of sulphide minerals (pyrite) by a group of acidophilic iron reducing micro organisms. Because of the salinisation properties of AMD and the associated scaling and bio-corrosion problems, as well as increased environmental awareness among the general population, methods are being investigated to remove the high sulphate concentration of AMD. Sulphate-rich effluents can be treated biologically when sulphate reducing bacteria (SRB) and organic matter are present. The products of the biological sulphate removal technology are sulphide and alkalinity, which contribute to the pH increase of the treated water. The advantage of the sulphide formation is that the metals present in the AMD precipitate as metal-sulphide (MeS). The disadvantage of the formation of hydrogen sulphide is its toxicity for both the methanogenic activity of granular sludge (Koster et al., 1986) as well for the sulphidogenic bacteria, e.g. *Desulfovibrio desulfuricans* (Okabe et al., 1995). Moreover, it has been reported that the produced sulphides are fatally toxic to humans at gaseous concentrations of 800 to 1 000 mg/l (Speece, 1996). Because of its toxicity, it is forbidden in most industrialised countries to drain sulphide-containing effluents into sewer pipes or surface waters (Janssen, 1996). When no metals are present, sulphide accumulation can result in a severe inhibition of the purification process and in some cases might even cause total process failure. Many studies have been dedicated to the effect of sulphide toxicity on the biological sulphate reduction efficiency. In gen-

eral, these studies demonstrated that, under mesophilic conditions, both granular and suspended sludge are more tolerant to sulphide inhibition at a pH of around 8, when the sulphide is in the HS⁻ form. At neutral pH values, free H₂S (sulphide in gas form), which is more toxic than HS⁻, accounts for 50% of total dissolved sulphide, whereas at pH 8 it is only around 10% (Lens and Hulshoff Pol, 2000). Speece (1996) listed the sulphide toxicity levels investigated by different researchers, which showed that 100 to 150 mg/l sulphide is toxic for lactate- and glucose-utilising SRB in a continuously stirred tank reactor (CSTR). Moreover, 60 to 75 mg/l sulphide was not tolerated by acetate- and propionate-utilising SRB (in a CSTR), while Parkin et al. (1990) reported that when the sulphide concentration was 60 to 70 mg/l, in an acetate- and propionate-fed chemostat, it resulted in process failure.

The aim of this investigation was to demonstrate that high sulphide concentrations (700 to 1 400 mg/l) in the biological sulphidogenic reactor can be beneficial for biological sulphate removal, operating:

- A completely mixed demonstration reactor
- A tall column reactor
- Four small-scale laboratory test reactors.

Materials and methods

Operating a pilot-scale completely-mixed biological sulphate removal reactor

Feed water

The feed water to the sulphate removing reactor consisted of AMD of which the composition is given in Table 1. It was supplemented with macronutrients (25 mg/l ammonia-N and

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Parameter	Composition AMD	Parameter	Composition AMD
pH	6.5	Magnesium (mg/l Mg)	101
Sulphate (mg/l SO ₄)	2 316	Iron (mg/l Fe)	2.5
COD (mg/l O ₂)	1 657	Aluminium (mg/l Al)	0.34
Chloride (mg/l Cl)	142	Manganese (mg/l Mn)	2.8
Sulphide (mg/l S)	0	Sodium (mg/l Na)	40
Alkalinity (mg/l) (CaCO ₃)	50	Total dissolved solids (mg/l)	4 117
Calcium (mg/l Ca)	957	Temperature (°C)	23

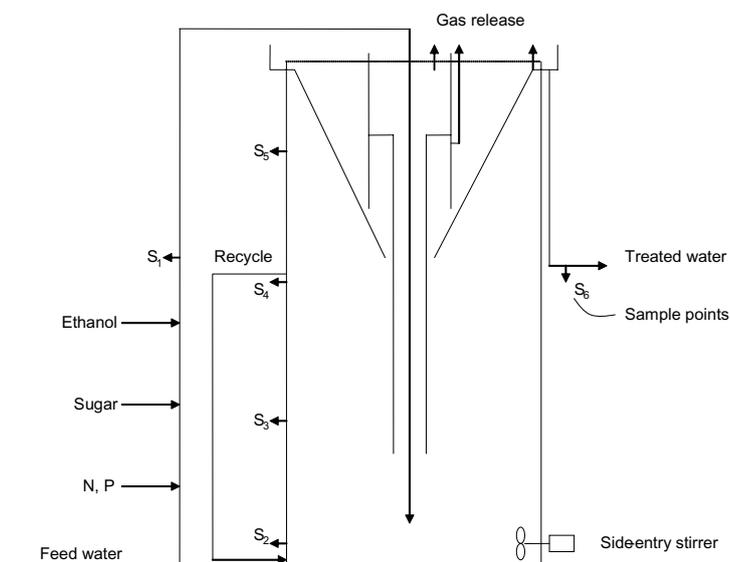


Figure 1
Schematic diagram of the anaerobic reactor

5 mg/l ortho-phosphate-P) and 3 mg/l Fe as the only micro-nutrient, since the feed (mine) water contained all trace elements required by the SRB. Sodium bicarbonate was added to maintain a pH of higher than 7.0. This procedure was terminated once sufficient alkalinity was generated from sulphate reduction. A mixture of ethanol (1.5 ml ethanol/l feed) and sugar (0.25 g sugar/l feed) was used as the carbon and energy source (Greiben et al., 2002). Sugar is degraded by the fermenting bacteria to volatile fatty acids (VFA) and hydrogen (H₂), which the SRB can use as the energy sources.

Reactor

The pilot-scale reactor system, CSIRosure, which was operated at Navigation Mine, Witbank, SA, consisted of a fibre-glass cylindrical reactor, with a built in cone shaped clarifier (Fig. 1). The volume of this reactor was 105.5 m³. The reactor was a completely mixed reactor system, fitted with a side-entry stirrer (260 r/min), positioned at the bottom of the reactor. The feed entered the reactor through a feed inlet pipe from the top to the bottom of the reactor and a recycle stream was in place from just above Sample Port S4 (Fig.1) to the bottom of the reactor. Complete mixing occurred in the reactor when the particles, entering the reactor through the feed pipe were dispersed immediately throughout the reactor due to the side entry stirrer as shown in Fig. 1. The reactor, which was operated at the ambient temperature of 20 to 25 °C, was inoculated with 10 m³ anaerobic digester sludge

from the Daspoort Sewage Treatment Plant, Pretoria. With time, this sludge became sulphate-adapted biomass with a large active population of SRB.

Experimental

CSIRosure

The experimental conditions of the continuously stirred reactor are given in Table 2.

Period (d)	Feed rate (m ³ /h)	HRT (h)
1-192	8	10.2
193-202	12	6.8
203-205	8	10.2
206-207	14	5.9
208-219	16	5.1

Operating a tall column reactor fitted with a venturi for optimal gas/liquid distribution for biological sulphate removal

Feed water

The artificial feed water, used for the venturi reactor, contained 3.9 g/l MgSO₄·7H₂O so that the SO₄ concentration was ≈1 700 mg/l. The nutrients P and N were added as 50 mg/l H₃PO₄ and 200 mg/l (NH₄)₂SO₄, respectively. Trace elements were added as 2 mg/l FeSO₄ and 1 ml/l of KCl, FeCl₃·4H₂O, NiNO₃·6H₂O, CoCl₂·6H₂O (all 1 mg/l) and MnCl₂·4H₂O, ZnCl₂, Na₂MoO₄, H₃BO₃ and CuCl₂·2H₂O (all 1 µg/l). The pH increased when additional sulphide was added to the feed water. In order to have the same start pH value in all three experiments, the feed-water pH was corrected using 1N HCl to 7.4 (Table 6).

Reactor

The venturi reactor (Fig. 2) was 6 m tall, its total volume was 190 l and its active volume was 157 l. A gas mixture of hydrogen and carbon dioxide (80% H₂; 20% CO₂) was used as the energy and carbon source. The gas mixture was sucked into the reactor through a venturi system, which provided small gas bubbles. The delivery rate of the gas was controlled using a Watson-Marlow pump (Watson-Marlow Bredel SA, Honeydew, Gauteng, SA). The feed water entered the reactor from the top, while the gas was pumped into the recycle line, thereby providing optimal conditions for gas/liquid exchange. The reactor was packed with 80% geo-textile strips for biofilm formation on these strips.

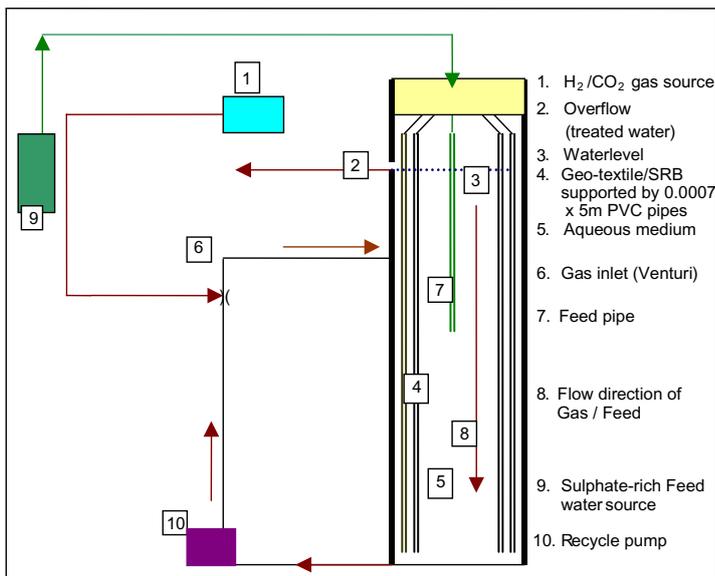


Figure 2
Schematic diagram of the venturi reactor

Experimental

The venturi reactor was operated in continuous mode until stable sulphate removal (90%) was achieved. Thereafter the reactor was operated in batch mode, as this way of operation provided conditions for collecting kinetic data in a short period of time, while investigating the effect of different parameters on the sulphate removal rate. Batch mode operation involved draining the treated water completely from reactor, so that only the biofilm-covered geo-textile strips stayed behind in the reactor, and subsequently re-filling the reactor with fresh feed water immediately thereafter. The gas mixture volume during this investigation was 700 ml/min (0.6 g/l H₂), resulting in a 4.5 times the stoichiometric amount needed for total sulphate reduction per litre of feed water (0.13 g/l). In order to investigate the effect of sulphide on the sulphate removal rate, three different studies were conducted, varying the sulphide concentration in the feed water. During the first study, no sulphide was added to the feed water (control), during the second study, the initial sulphide concentration in the feed water was 100 mg/l, while in the third study this was 268 mg/l. In both instances Na₂S was used (Merck SA, Johannesburg). Samples (50 ml) were taken

at regular intervals over a period of 24 h during each investigation.

Operating four stirred laboratory batch-test reactors

Feed water and reactors

The feed water for four stirred batch-test reactors (1 to 4), consisted of artificial feed water of which the sulphate concentration (MgSO₄) was approximately 1 500 mg/l. The feed water was supplemented with both macro-nutrients (75 mg/l ammonia-N and 15 mg/l orthophosphate-P) and micro-nutrients (100 µg/l Fe, 210 µg/l Co, 0.28 µg/l Mn, 0.44 µg/l V, 0.25 µg/l Ni, 0.48 µg/l Zn, 0.40 µg/l Mo, 0.18 µg/l B, 0.37 µg/l Cu). The carbon and energy source added to the feed water was a mixture of ethanol (1.5 ml/l) and sugar (0.25 g/l), such that the COD: SO₄ ratio was 1:1. Glass bottle reactors 1 to 4 (volume 2 l) were operated under anaerobic conditions. The reactors received sulphate-adapted biomass, obtained from the CSIRosure reactor. Samples (50 ml) were taken daily from the bottom of the reactor. The experiments were carried out at room temperature (20 to 25°C).

Experimental

The experimental conditions operating Reactors 1 to 4 are given in Table 3. The following information can be observed from Table 3:

- **Day 1:** The conditions in all 4 reactors were the same.
- **Day 2:** After sample-taking, the contents of Reactors 1 to 3 were disposed of and these reactors received new biomass (250 ml) and 1 750 ml fresh feed water. Therefore the conditions in Reactors 1 to 3 on Day 2 were identical to those in Reactors 1 to 4 on Day 1. Reactor 4: After sample-taking (50 ml), fresh MgSO₄ as a powder was dissolved in 50 ml tap water (to replace the volume taken for the sample) and added to the contents of Reactor 4, such that the start SO₄ concentration on Day 2 was 1 500 mg/l. Sulphide-rich water was produced in Reactor 4 on Day 1, which remained in the reactor. Therefore on Day 2, Reactor 4 contained an increased sulphide concentration compared to Reactors 1 to 3.
- **Day 3:** The procedures as described for Day 2 were repeated for Reactors 1 and 2, so that the conditions in Reactors 1 and 2 were the same and identical to all four reactors on

TABLE 3				
The experimental conditions for the batch test				
Period	Reactors			
	1	2	3	4
Day 1	1 750 ml feed 250 ml biomass	1 750 ml feed 250 ml biomass	1 750 ml feed 250 ml biomass	1 750 ml feed 250 ml biomass
Day 2	1 750 ml feed 250 ml biomass**	1 750 ml feed 250 ml biomass **	1 750 ml feed 250 ml biomass**	Fresh SO ₄ feed*
Day 3	1 750 ml feed 250 ml biomass**	1 750 ml feed 250 ml biomass**	Fresh SO ₄ feed*	Fresh SO ₄ feed*
Day 4	1 750 ml feed 250 ml biomass**	Fresh SO ₄ feed*	Fresh SO ₄ feed*	Fresh SO ₄ feed*
Day 5	Fresh SO ₄ feed*	Fresh SO ₄ feed*	Fresh SO ₄ feed*	Fresh SO ₄ feed*

* Fresh SO₄ (Mg SO₄ as chemical :1 500 mg/l SO₄) was added to the contents of reactor
 ** Reactors received fresh biomass, fresh feed and nutrients

Day	1		2		3		4	
	VSS	pH	VSS	pH	VSS	pH	VSS	pH
1	4 004	7.82	3 268	8.31	4 044	8.01	3 788	7.64
2	4 860	8.12	5 316	8.09	5 124	8.08	3 408	7.59
3	6 908	8.23	3 636	8.16	4 652	7.75	3 268	7.55
4	7 788	8	4 488	7.94	3 652	7.83	2 916	7.44
5	4 892	8.16	2 656	7.96	2 720	7.69	2 320	7.38

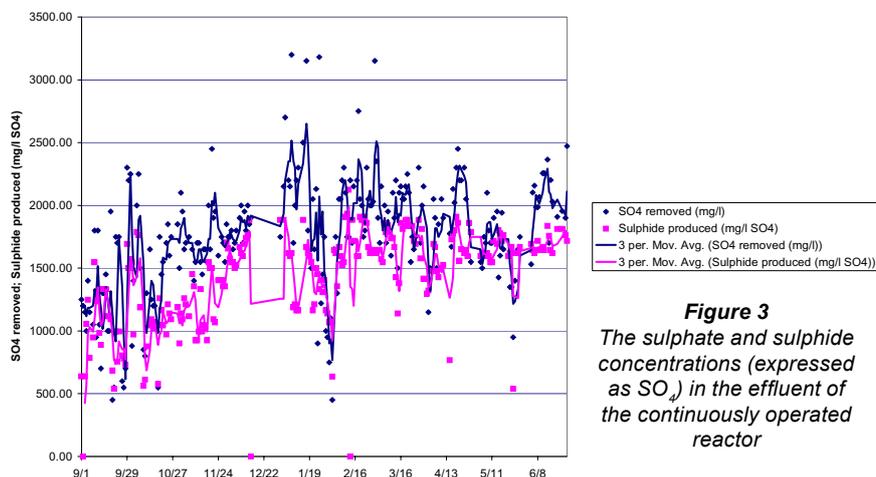


Figure 3
The sulphate and sulphide concentrations (expressed as SO₄) in the effluent of the continuously operated reactor

were carried out according to standard analytical procedures as described in *Standard Methods*, 1985. The redox potential of the samples was calculated from the mV and stabilisation temperature measured with a 744 pH meter (Metrohm). With the exception of the MLSS, VSS, sulphide and feed COD, all analyses were carried out on filtered samples (Whatman #1). The COD samples were pretreated to eliminate the sulphide contribution to the COD concentration.

Results and discussion

Sulphate removal/sulphide production

Operating the CSIRosure pilot plant reactor for biological sulphate removal

Sulphate reduction was obtained during the operation of the demonstration plant. When the hydraulic retention time (HRT) was 5.1 h, the SO₄ removal rate was 12 g SO₄/ℓ·d. The sulphate reduction (mg/ℓ SO₄) and the sulphide production (expressed as mg/ℓ SO₄), measured over a period of 9 months in the continuous reactor, is given in Fig. 3. The graphs in Fig. 3 show the relationship between sulphate removed and sulphide produced in the anaerobic reactor (Eq. (1)). The sulphide concentration in the effluent was stoichiometrically slightly less than the removed sulphate concentration:



The fact that slightly less sulphide was produced than sulphate reduced, as shown in Fig. 3 can be ascribed to the biological sulphide oxidation to sulphur, due to air diffusion on top of the clarifier, which showed a white layer of produced sulphur (Eq. (2)).



Moreover, metal sulphide precipitation occurred due to the metals concentration in the AMD (Eq. (3)):



During continuous sulphate reduction, high levels of sulphide, often higher than 600 mg/ℓ, were produced (Fig. 3). At these high sulphide concentrations, sulphate removal continued, indicating that the SRB could proceed with sulphate removal despite the high sulphide concentration in the reactor. This

Day 1. The procedures as described for Reactor 4 on Day 2 were repeated for Reactors 3 and 4 on Day 3 and thus Reactor 3 had sulphide rich water, while Reactor 4 contained higher concentration sulphide- rich water.

- **Day 4:** The same procedures as for Reactors 1 to 3 on Day 2 were repeated for Reactor 1, thus the conditions in Reactor 1 were the same as those in all four reactors on Day 1. The same procedure as for Reactor 4 was repeated for Reactors 2, 3 and 4, therefore these reactors contained sulphide-rich water: The sulphide concentration in R 4 > in R 3 > in R 2.
- **Day 5:** All 4 reactors were similarly treated as Reactor 4 on Day 2 and thus all reactors received the same fresh SO₄ feed (MgSO₄ as the chemical, dissolved in 50 ml water) and all 4 reactors contained sulphide-rich feed: The sulphide concentration in R 4 > in R 3 > in R 2 > in R 1.

The reactor pH and the VSS concentrations during the batch test are given in Table 4. It can be observed that the reactor pH mainly ranged between pH of 7.5 to 8.5.

Although the conditions in all four reactors were the same on Day1, it can be noted that the VSS values differed slightly, varying from 3.2 to 4.0 g/ℓ. From Day 2 to 5, the VSS concentration in R4 decreased slowly, because biomass was lost during sample taking while no new biomass was added. On Day 2 the VSS values in Reactors 1 to 3 were similar. From Day 2 onwards the VSS concentration decreased in R3, for the same reason as in R4. The lower VSS concentration in the reactors will result in lower volumetric sulphate reduction reaction rates, while the specific sulphate reduction rate is a function of the VSS (biomass) concentration.

Analytical

Determinations of sulphate, sulphide, COD, pH, mixed liquor suspended solids (MLSS) and volatile suspended solids (VSS)

result is contrary to the findings of McCartney and Oleszkiewicz (1991; 1993), Postgate (1984), Klemp et al. (1985), Shimada (1987), Hilton and Oleszkiewicz, (1988) and Okabe et al. (1992) who found that sulphate reduction could not continue, due to sulphide concentrations higher than at the most 500 mg/l (as S). However, most of those studies were operated at a reactor pH of 6.5 to 7.0, whereas the reactor pH in this study was maintained at an average pH of 7.5 (Fig. 4). Sulphide toxicity is strongly dependent on pH, because the chemical equilibrium of sulphide is pH dependent. At pH 8 most of the total sulphide is in the HS⁻ form. Molecular H₂S (g) has been found to be the major toxic form of sulphide because H₂S (g) can pass through the lipid cell membrane into the cytoplasm (Speece, 1996). The exact sulphide inhibition mechanisms have not been explained yet. Many studies have been dedicated to sulphide toxicity on sulphate reduction. The results of these toxicity studies have been described by several researchers and have been listed in Table 5 (adopted from Lens et al., 1998). In general, these studies demonstrated that, under mesophilic conditions, both granular and suspended sludge are more resistant to sulphide inhibition at a higher pH of around 8. This finding can be ascribed to the fact that sulphide is dominant at pH of 6.5. It was also shown that the sludge was more sensitive to the concentration of H₂S (g) than to the concentration of total sulphide (TS). Weast (1981) described that the pK_a value of the dissociation equilibrium of H₂S is about 7.04 at 18 °C. Above pH 8.0 to 9.0 virtually all dissolved sulphide is present in its ionised form, while at neutral pH values 20 to 50% of the dissolved sulphide is present as H₂S (g), depending on the reactor temperature (O'Flaherty and Collieran, 2000).

Operating the venturi reactor using hydrogen gas as the energy source

Figure 5 shows the average sulphate removal when the initial sulphide concentrations in the feed water were 0, 100 and 268 mg/l, while Table 6 shows the overall results of the three experiments.

It can be seen (Table 6) that the maximum sulphate removal (1 475

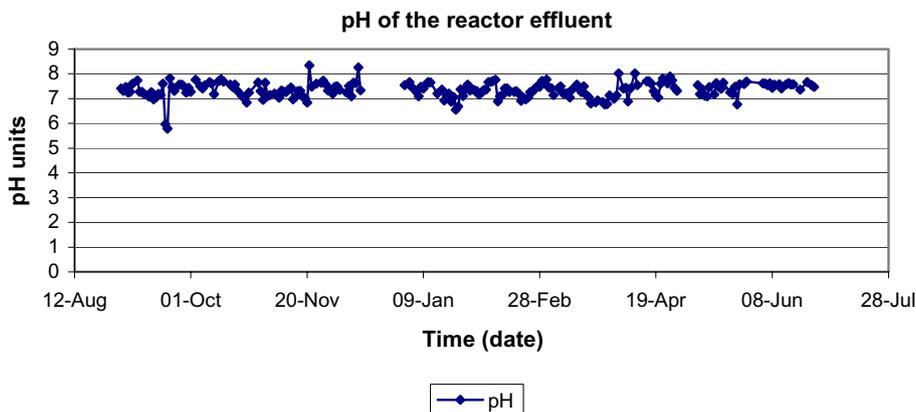


Figure 4
The pH in the reactor effluent

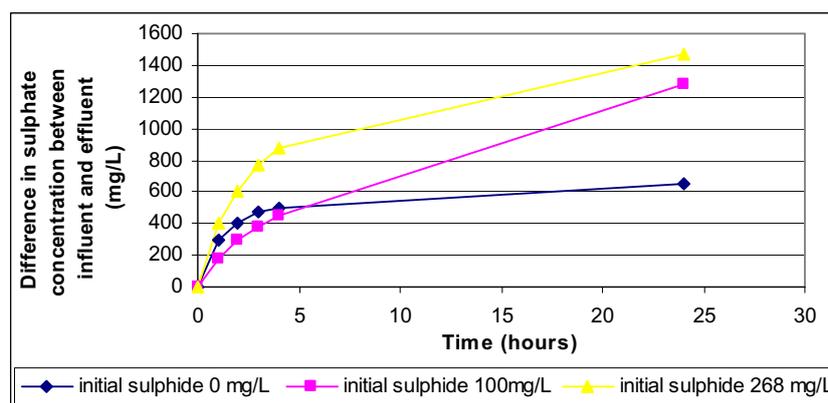


Figure 5
The average sulphate removal operating the venturi reactor at three different initial sulphide concentrations (0, 100 and 268 mg/l)

Sludge type	Substrate	T (°C)	pH	H ₂ S (mg/l)	TS (mg/l)	Reference
<i>Desulfovibrio desulfuricans</i>	Lactate	35	7.0	250	500	Okabe et al. (1992)
Sludge susp.	Lactate/Acetate	35	7.2-7.6	NR	83	McCartney and Oleszkiewicz (1991)
Sludge susp.	Lactate	35	7.0	>300	NR	McCartney and Oleszkiewicz (1993)
Sludge granules	Acetate	30	7.0 -7.4	171	615	Visser et al. (1996)

* sludge susp. = sludge suspension ** NR= Not reported

mg/l) was obtained when the initial feed sulphide concentration was 268 mg/l, which increased to 500 mg/l after 24 h of batch operation. Similar results could be observed from the study in which the initial sulphide concentration was 100 mg/l, resulting in the maximum sulphate removal of 1 275 mg/l and the final sulphide concentration of 452 mg/l. When no sulphide was added to the feed water, the maximum sulphate removal was only half the amount (650 mg/l) of the results in the studies when sulphide was added to the feed water. These results show that due to the addition of sulphide to the feed water, the sulphate removal rate increased. When calculating the sulphide_{produced}/sulphate_{removed} ratio, it can be seen from the data in Table 6 that when 0, 100 and 268 mg/l sulphide was added to the feed, this ratio varied from 0.36 to 0.27 and 0.16, respectively, while the theoretical value is 0.33 (S²⁻/SO₄²⁻ = 0.33). It can be hypothesised that due to the increased sulphide concentration in the feed, improved sulphate reduction occurred resulting in the formation of intermediate products, such

Parameter	Sulphide 0 mg/l	Sulphide 100 mg/l	Sulphide 268 mg/l
Initial feed S ²⁻ (mg/l)	0	100	268
Final reactor S ²⁻ (mg/l)	236	452	500
SO ₄ removed over 24 h (mg/l)	650	1 275	1 475
S ²⁻ _{produced} /SO ₄ _{reduced} ratio	0.36	0.27	0.16
Start pH	7.4	7.4	7.4
Final pH	7.5	7.8	7.86
Start redox potential	+175	-164	-169
Final redox potential	-148	-159	-187.5

reactor. As mentioned before, a reactor pH > 7.5 causes a significant reduction of the dissolved H₂S concentration and results in a shift toward the dissociated HS⁻ ion.

Operating four stirred reactors under batch conditions

The sulphide concentrations in reactors R1 to 4 are given in Fig. 7. It can be seen that the sulphide concentration on Day 1 was the same in all 4 reactors, but that from Day 2 it started to increase in Reactor R4. After Day 2, it increased in both R4 and R3, where after (Days 4 and 5) it increased in R2 and R1 as well. After the Day 3 (72 h), it had increased in all four reactors, with a maximum at 96 h (4 d), when the sulphide concentrations were

1424, 1616, 1424 and 824 mg/l, in R4 to R1, respectively. At time 98 h, all reactors showed an initial decrease in the sulphide concentration, but during the course of Day 5, the concentration increased again. The decrease in the sulphide concentration can possibly be ascribed to the fact that the reactor bottles were opened at the top and that air entered the reactor flasks, causing biological sulphide oxidation. Some sulphur-forming could be observed at the top of the reactors.

The volumetric and specific sulphate reduction rates were calculated, which are given in Table 7. The results show that the specific SO₄ reduction rates increased from 0.61 to 0.99 to 1.49 g SO₄/gVSS, when the sulphide concentration increased from 132 to 644 to 1 424 mg/l in R4. Similar results were obtained in the other three reactors: when the sulphide concentration increased, the sulphate reduction rates increased. It can be observed that the high sulphide concentrations did not inhibit the sulphate reduction rates, but on the contrary seemed to stimulate the sulphate reduction rates. This observation is contrary to the findings of Reis et al. (1992); however, they maintained their reactors at a pH of 6.2 to 6.4. It can be speculated that at the operating reactor temperature of 20 to 25°C and at a reactor pH of approximately 8, the methanogenic bacteria (MB) were inhibited by the higher sulphide concentrations, thus providing the SRB with an opportunity to utilise the available substrate. It has also been reported (Lens and Hulshoff Pol, 2000) that using sulphate-adapted biomass will favour the substrate utilisation by SRB over MB as was the case in this study.

Visser (1995) showed that the substrate utilisation rate increased at higher sulphide concentrations, showing that growth and activity of the SRB were uncoupled. The study of Okabe et al. (1992) showed that when the sulphide concentration increased, the cell size of the SRB (*Desulfovibrio desulfuricans*) decreased, which can partly explain the decreased cell yield at increasing sulphide concentrations during the batch study, namely, when the sulphide concentration increased, the VSS concentration decreased (Table 4). The VSS concentration decrease likely caused the decrease in the sulphate reduction rates as shown in Tables 4 and 7.

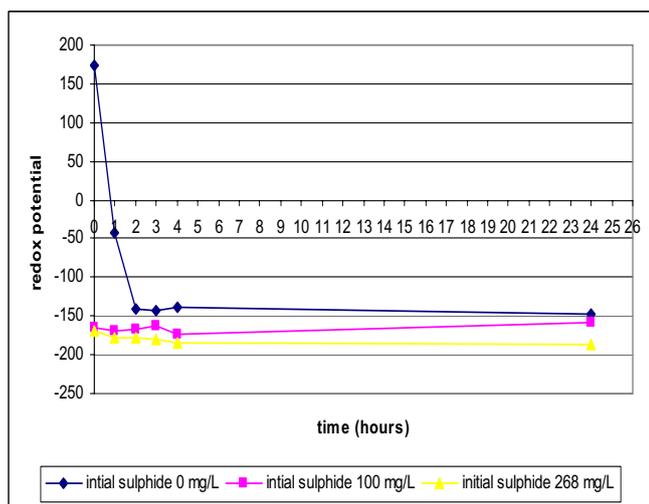


Figure 6
Effect of initial sulphide concentration on redox

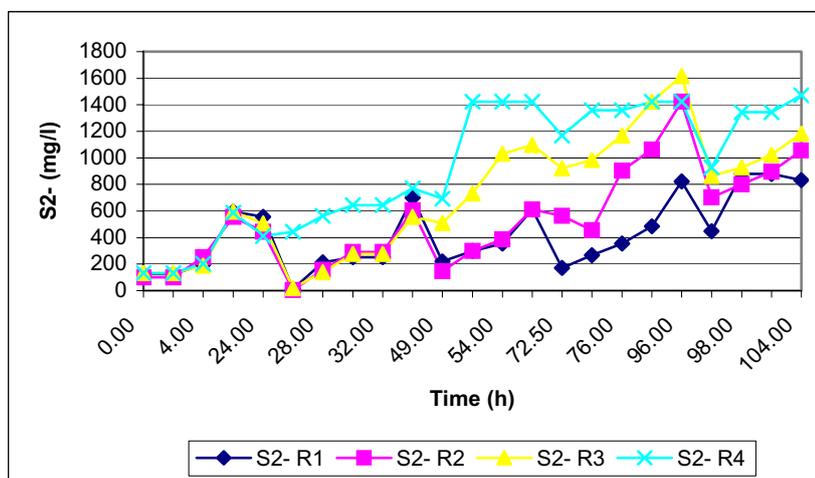


Figure 7
The sulphide concentration in Reactors 1 to 4

a thiosulphate and sulphite, rather than H₂S. The reduction of SO₄²⁻ to H₂S proceeds through a number of intermediate stages (Brock, 1997). The data in Table 6 show that during the three experiments the reactor pH increased and that the redox potential decreased (Fig. 6). These results show that improved sulphate reduction is obtained when the reactor sulphide concentration and pH increased and when the redox potential decreased in the

	R 1		R 2		R 3		R 4	
Day	Vol R.R* g SO ₄ /ℓ	Spec. R.R** g SO ₄ /gVSS	Vol R.R* g SO ₄ /ℓ	Spec. R.R** g SO ₄ /gVSS	Vol R.R* g SO ₄ /ℓ	Spec. R.R** g SO ₄ /gVSS	Vol R.R* g SO ₄ /ℓ	Spec. R.R** g SO ₄ /gVSS
1	2.26	0.56	1.94	0.59	1.92	0.47	2.30	0.61
2	2.64	0.54	2.64	0.5	2.52	0.49	3.36	0.99
3	2.40	0.35	2.40	0.43	4.68	1.01	4.87	1.49
4	2.86	0.37	3.43	0.76	2.64	0.72	3.07	1.05
5	3.82	0.78	3.12	1.17	2.64	0.97	2.62	1.13

The treatment of sulphide-rich effluents

Although the production of sulphide after sulphate reduction might enhance the sulphate removal process, sulphide in aqueous and gaseous solution can cause chemical problems, such as corrosion, odour and an increase of the effluent COD, as sulphide oxidises when in contact with air. Several methods have been described to remove sulphide from the effluents, such as the partial biological sulphide oxidation (Buisman, 1989; Janssen, 1996) to elemental sulphur (S⁰). This is a relatively inexpensive option, especially when partial sulphide oxidation can be achieved in the same reactor in which the sulphate reduction is accomplished (Maree et al., 1997). Colourless sulphur bacteria (*Thiobacillus* spp.) oxidise sulphide to S⁰ or to sulphate in the presence of air (O₂). Chemically, sulphide can be removed through stripping with CO₂ gas, after which the gas mixture is absorbed with a ferric (Fe³⁺) solution, producing Fe²⁺ and elemental sulphur (Maree et al., 2004). The produced Fe²⁺ can be oxidised to Fe³⁺ to maintain the reaction.

Results summarised

The results of the studies operating the continuous reactor, the venturi reactor as well as the batch-test reactors indicated that high sulphide concentrations (varying from 500 to 1 500 TS mg/ℓ) in an anaerobic sulphidogenic bioreactor, operating at a reactor pH range of 7.5 to 8.5 and a temperature range of 20 to 25°C did not inhibit the sulphate reduction rates. When the sulphide concentration was 700 mg/ℓ in the continuous reactor, the sulphate reduction rate was 12 g SO₄/ℓ·d. It was shown that when sulphide was added to the reactor feed water (100 and 268 mg/ℓ, respectively), operating a tall venture reactor, using hydrogen as the energy source, the maximal sulphate removal was 1 275 and 1 475 mg/ℓ, while when no sulphide was added to the feed water, the maximal sulphate removal was 650 mg/ℓ. The results of the batch tests showed that when the sulphide concentration in the bio reactor increased, the sulphate reduction rate increased as well. When the sulphide concentration was 1 400 mg/ℓ in the batch test, the volumetric and specific sulphate reduction rates were 4.9 g SO₄/ℓ·d and 1.5 g SO₄/g VSS, respectively. It was speculated that the high sulphide concentration resulted in failure of the MB performance, thus favouring SRB for the substrate competition. The increased sulphate reduction rates observed when the sulphide concentrations increased may be due to an increased thermodynamic driving force for sulphate reduction because of the low redox potential in the reactor. However, the apparent increase in sulphate reduction subsequently as sulphide concentrations increased may not be because of increased thermodynamic driving force. This is because the system is closer to equilibrium position. This relative "distance" from the equilibrium position is not likely to change

because an increase in pH is theoretically expected to result in a decrease in redox potential when a dominant redox couple, like sulphate: sulphide is present. This might suggest that the reason for increased sulphate reduction is more a kinetic than a thermodynamic one. Since the results of the studies described in this paper are based on observations, further research needs to be conducted to better understand the exact mechanisms why the produced sulphide at a reactor pH >7.5 is stimulatory for the biological sulphate reduction.

Conclusions

When operating three different reactor configurations, applying the biological sulphate reduction technology, the produced sulphide did not inhibit the biological sulphate removal process. Operating a complete mixed pilot scale reactor resulted in a sulphate removal of 12 gSO₄/(ℓ·d), while the average sulphide production was 600 mg/ℓ. Under batch test operation, the increased sulphide concentration (132, 644 and 1424 mg/ℓ) resulted in improved specific sulphate removal rates (0.61, 0.99 and 1.49 gSO₄/gVSS). When operating the Column Venturi reactor and sulphide was added to the feed water (100 and 368 mg/ℓ Na₂S), the sulphate removal was 1 275 and 1 475 mg/ℓ, respectively. This compared favourably with the sulphate removal of 650 mg/ℓ, when no sulphide was added to the feed water. The following trend was observed: good sulphate reduction resulted in high sulphide and alkalinity production, which caused a substantial increase in the reactor pH. When operating the column reactor a decrease in the reactor redox potential was observed. The high concentration of sulphide can be removed by the biological sulphide oxidation process or by the chemical stripping technology. In both cases elemental sulphur is the end product.

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